**Quick No Calc Gas Laws**

1. A 8.0L sample of gas is at pressure of 1.25atm. If the pressure increased to 5.0atm, what is the new volume?
2. A sample at -23oC occupies a volume of 25.0mL. If the new volume is 37.5mL, find the new temperature assuming pressure is held constant.
3. 8.21L of a gas at 27oC is at 1.0atm. Find the number of moles.
4. 3 moles of gas occupies 225mL of space. If pressure and temperature are held constant but an additional 1 mole of gas is added, what is the new volume?
5. Ethane combusts in the presence of oxygen according to the following reaction:

2C2H6 + 7O2(g) → 4CO2 (g) + 6 H2O (g)

If 6.0mol of ethane burn in the presence of 24.0mol O2(g), how many moles of each product form? How many moles of gas are in the container when the reaction ends?

What is the mole fraction of the CO2 at the end of the reaction?

**Quick No Calc Gas Laws**

1. A 8.0L sample of gas is at pressure of 1.25atm. If the pressure increased to 5.0atm, what is the new volume?
2. A sample at -23oC occupies a volume of 150.0mL. If the new volume is 37.5mL, find the new temperature assuming pressure is held constant.
3. 8.21L of a gas at 27oC is at 1.0atm. Find the number of moles.
4. 3 moles of gas occupies 225mL of space. If pressure and temperature are held constant but an additional 1 mole of gas is added, what is the new volume?
5. Ethane combusts in the presence of oxygen according to the following reaction:

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