**Dougherty Valley HS AP Chemistry**

**S-31**

**Equilibrium Review**

**Quick Check #2**

**Name: Date: Period: Seat #:**

Check off each item if you can do the question. Write down a question to ask if you cannot do the question.

* **Reaction Quotient**

 H2(g) + Br2(g)  2 HBr(g) Kc = 5.5 x 103

 [H2]=0.10 M [Br2]=0.20 M [HBr]=8.5 M

What will happen to the [HBr] as this reaction

approaches equilibrium? (Show your calculation.)

* **Kp & Kc**

 2 NO(g) + Br2(g)  2 NOBr(g)

 Kc = 1.2 x 10-10 at 25 °C

Write the Kp expression for this reaction and

calculate its value. [R = 0.0821 L·atm/mol·K]

* **Le Châtelier’s’ Principle Demo**

 Co(H2O)62+(aq) + 4 Cl(aq)  CoCl42(aq) + 6 H2O(l)

 pink blue

 a) add HCl(aq) \_\_\_\_ \_\_\_\_ \_\_\_\_ \_\_\_\_

 b) add H2O(l) \_\_\_\_ \_\_\_\_ \_\_\_\_ \_\_\_\_

 c) increase the temperature \_\_\_\_ \_\_\_\_ \_\_\_\_ \_\_\_\_

 d) decrease the temperature \_\_\_\_ \_\_\_\_ \_\_\_\_ \_\_\_\_

 e) add AgNO3(aq) \_\_\_\_ \_\_\_\_ \_\_\_\_ \_\_\_\_

Note:

 Predict (a) and (b) before the demonstration.

 Watch (c) and determine whether the reaction is endo- or exo-thermic.

 Predict (e) before the demonstration.