

Name: _____

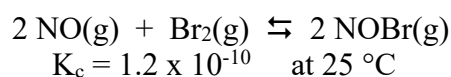
Date: _____

Period: _____

Seat #: _____

Check off each item if you can do the question. Write down a question to ask if you cannot do the question.

K_p & K_c



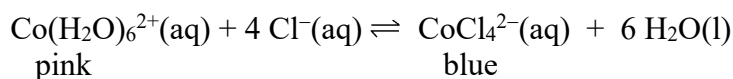
Write the K_p expression for this reaction and calculate its value. [R = 0.0821 L·atm/mol·K]

ICE Box Problem

A solution is prepared by dissolving 0.050 mol of diiodocyclohexane, C₆H₁₀I₂, in the solvent CCl₄. The total solution volume is 1.00 L. When the reaction, C₆H₁₀I₂ ⇌ C₆H₁₀ + I₂, comes to equilibrium, the concentration of I₂ is 0.035 mol/L. What are the concentrations of C₆H₁₀I₂ and C₆H₁₀ at equilibrium?

C ₆ H ₁₀ I ₂	⇌	C ₆ H ₁₀	+	I ₂

Le Châtelier's' Principle Demo



- | | | | | |
|-------------------------------|-------|-------|-------|-------|
| a) add HCl(aq) | _____ | _____ | _____ | _____ |
| b) add H ₂ O(l) | _____ | _____ | _____ | _____ |
| c) increase the temperature | _____ | _____ | _____ | _____ |
| d) decrease the temperature | _____ | _____ | _____ | _____ |
| e) add AgNO ₃ (aq) | _____ | _____ | _____ | _____ |

Note:

Predict (a) and (b) before the demonstration.

Watch (c) and determine whether the reaction is endo- or exo-thermic.

Predict (e) before the demonstration.