**Dougherty Valley HS AP Chemistry**

**S-37**

**Atomic Structure**

**Quick Check #1**

**Name: Date: Period: Seat #:**

|  |
| --- |
| Formulas and Constants |
| c = υ | υ =  |  =  | E = hυ | E =  |
| c = 2.998 x 108 m/s h = 6.626 x 10-34 J·s |

🞎 List all electromagnetic radiations from low energy to high.

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
|  |  |  | **R O Y G B V** |  |  |  |

🞎 We can see electromagnetic radiation with wavelengths between 400 nm and 700 nm.

 Is 400 nm red light or violet light? \_\_\_\_\_\_\_\_\_\_\_\_\_\_ Justify your answer.

🞎 Consider this graphic from the **Online Tutorial: Electromagnetic Radiation**. The scale is in nm.



 What is the wavelength (λ) of this wave? \_\_\_\_\_\_\_nm

 Would you be able to see this wave? \_\_\_\_\_\_

 What is this wavelength in meters? \_\_\_\_\_\_\_\_\_\_\_m

🞎 Yellow light from a sodium vapor light has a wavelength of 589 nm. Calculate the frequency of this color of yellow light in Hz.

🞎 A radio station (KPCC) has a frequency of 89.3 MHz (megahertz).

 How many Hz are in a MHz? \_\_\_\_\_\_\_\_\_

 What is the frequency of this radio wave in s-1? \_\_\_\_\_\_\_\_

 What is the energy of the radio waves being emitted (in Joules)?