

Name: _____ Date: _____ Period: _____ Seat #: _____

Try these problems. If you can DO them, check the box (). If you CANNOT do them, write some notes TO YOURSELF about what you need to study to succeed at these problems.

Formulas: Quickly write the formulas for the following concentration units:

Molality	Weight Percent	Mole Fraction	Molarity

Dissecting A Given Concentration:

The concentration of a NaOH solution is 0.25 m. This translates into 0.25 and 1.0.

0.25 = _____ and 1.0 = _____

The concentration of a HC₂H₃O₂ solution is 5.00% by weight. This translates into 5.00 and 100.

5.00 = _____ and 100 = _____ and 95.0 = _____

Change one concentration into another:

Household vinegar is labeled as 5.00% by weight. It has a density of 1.01 g/mL. Fill in the chart.

	mass (grams)	moles (mol)	volume (L)
solute			
solvent			
solution			



Molality	Mole Fraction	Molarity