**Dougherty Valley HS AP Chemistry**

**S-69**

**Solutions**

**Quick Check**

**Name: Date: Period: Seat #:**

* **Ornament Chemistry:** Write the balanced chemical equation for the following. Indicate the oxidation numbers of the elements.

Zn(s) + HCl(aq) = \_\_\_\_\_\_\_ + \_\_\_\_\_\_\_

|  |  |  |
| --- | --- | --- |
|  | **Reactants** | **Products** |
| Zn |  |  |
| H |  |  |
| Cl |  |  |

Fe(s) + CuSO4(aq) = \_\_\_\_\_\_\_ + \_\_\_\_\_\_\_

|  |  |  |
| --- | --- | --- |
|  | **Reactants** | **Products** |
| Fe |  |  |
| Cu |  |  |
| S |  |  |
| O |  |  |

* **Terminology:**  
  In the first reaction, \_\_\_\_\_\_\_\_\_\_ is getting oxidized. \_\_\_\_\_\_\_\_ is the oxidizing agent.  
  In the first reaction, \_\_\_\_\_\_\_\_\_\_ is getting reduced. \_\_\_\_\_\_\_\_ is the reducing agent.

In the second reaction, \_\_\_\_\_\_\_\_\_\_ is getting oxidized. \_\_\_\_\_\_\_\_ is the oxidizing agent.  
 In the second reaction, \_\_\_\_\_\_\_\_\_\_ is getting reduced. \_\_\_\_\_\_\_\_ is the reducing agent.