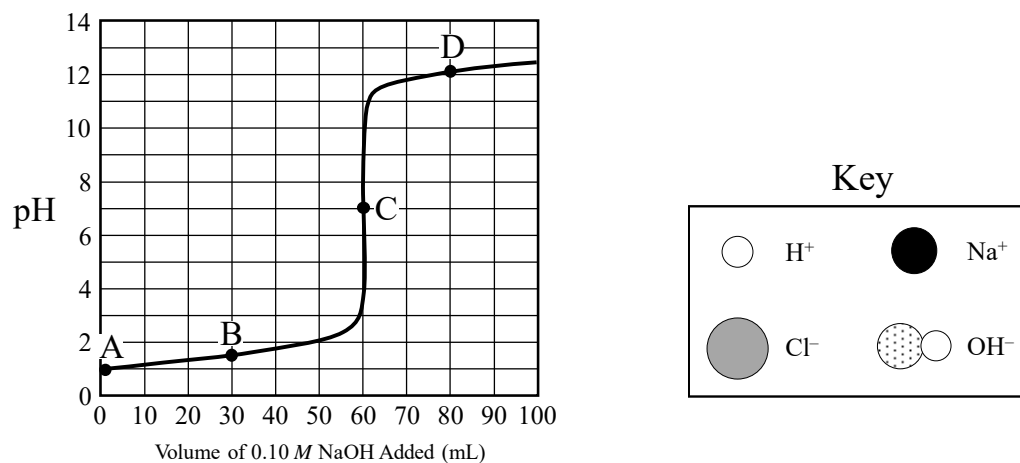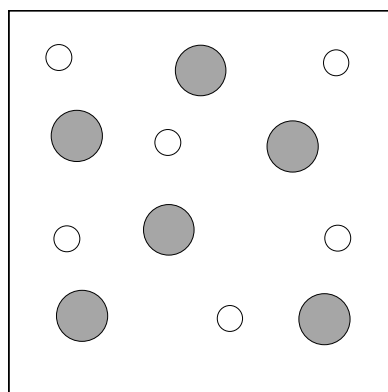A 60.0 mL sample of 0.10 M HCl(aq) is titrated with 0.10 M NaOH(aq). The pH of the resulting solution is measured with a pH meter and graphed as a function of the volume of the titrant added.

Key

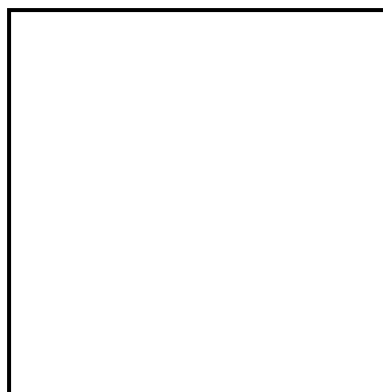Volume of 0.10 M NaOH Added (mL)

The diagram labeled as "A" below represents a particulate representation of a small representative portion of the solution at point A of the pH curve shown above. For clarity, water molecules are not shown.
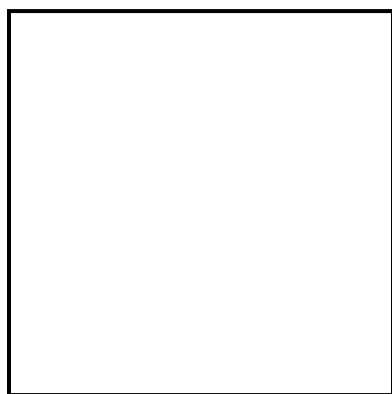
Use the symbols shown in the Key above to draw particulate representations of the solution at points B, C, and D of the pH curve shown above.
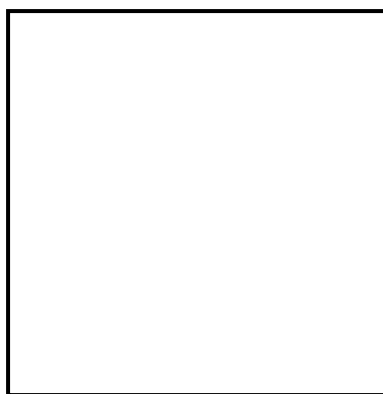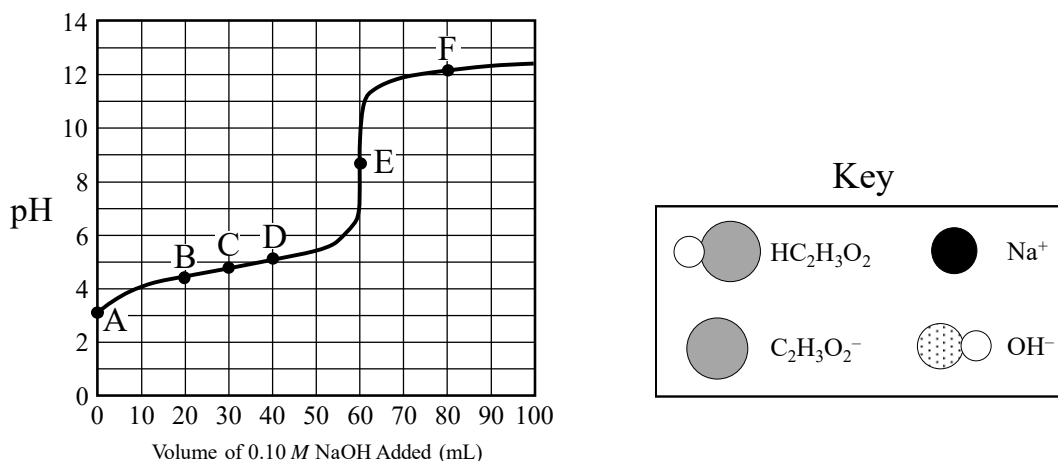

A

B

C

D

A 60.0 mL sample of 0.10 M $HC_2H_3O_2(aq)$ is titrated with 0.10 M $NaOH(aq)$. The pH of the resulting solution is measured with a pH meter and graphed as a function of the volume of the titrant added.

The diagram labeled as "A" below represents a particulate representation of a small representative portion of the solution at point A of the pH curve shown above. For clarity, water molecules are not shown.

Use the symbols shown in the Key above to draw particulate representations of the solution at points B, C, D, E, and F of the pH curve shown above.
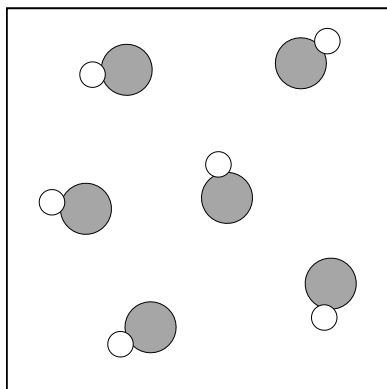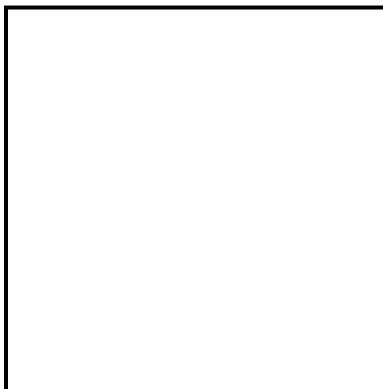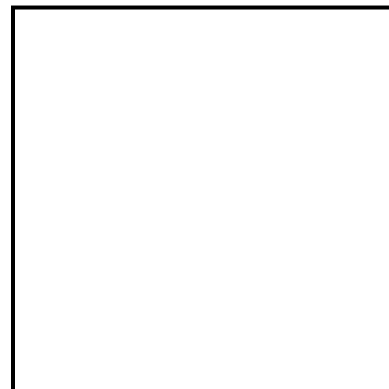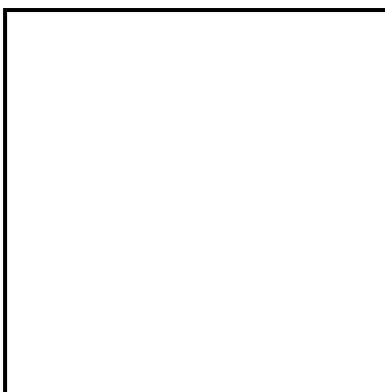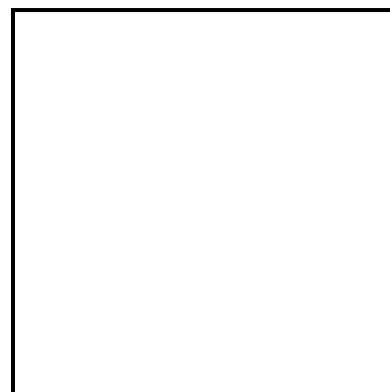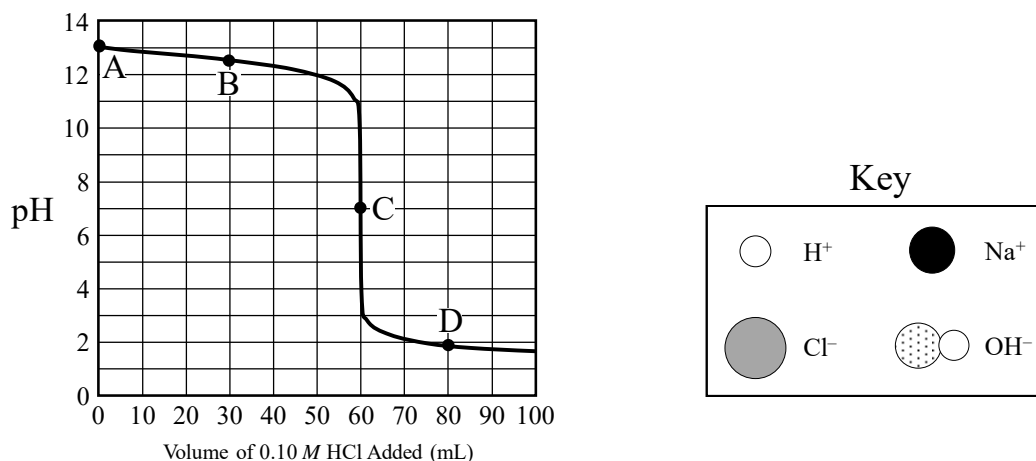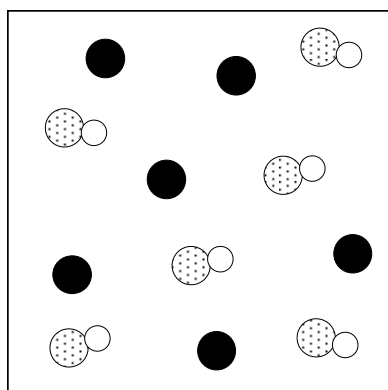

A

B

C

D

E

F

A 60.0 mL sample of 0.10 M NaOH(aq) is titrated with 0.10 M HCl(aq). The pH of the resulting solution is measured with a pH meter and graphed as a function of the volume of the titrant added.

pH (y-axis, values 0, 2, 4, 6, 8, 10, 12, 14)

Points labeled A, B, C, D on the curve.

Volume of 0.10 M HCl Added (mL) (x-axis, values 0 10 20 30 40 50 60 70 80 90 100)

Key

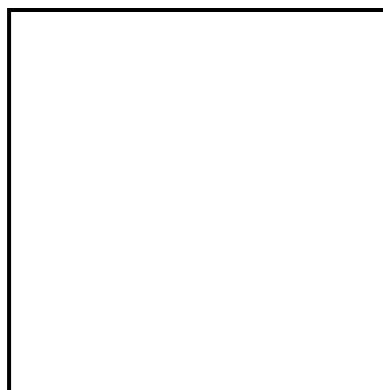○ H^+ ● Na^+

◔ Cl^- ◌○ OH^-

The diagram labeled as "A" below represents a particulate representation of a small representative portion of the solution at point A of the pH curve shown above. For clarity, water molecules are not shown.

Use the symbols shown in the Key above to draw particulate representations of the solution at points B, C, and D of the pH curve shown above.
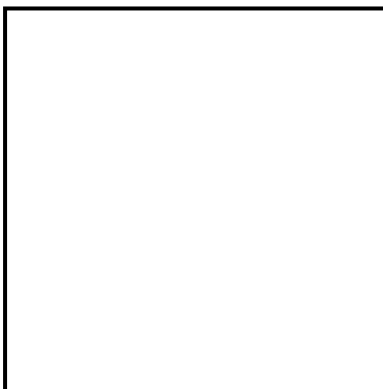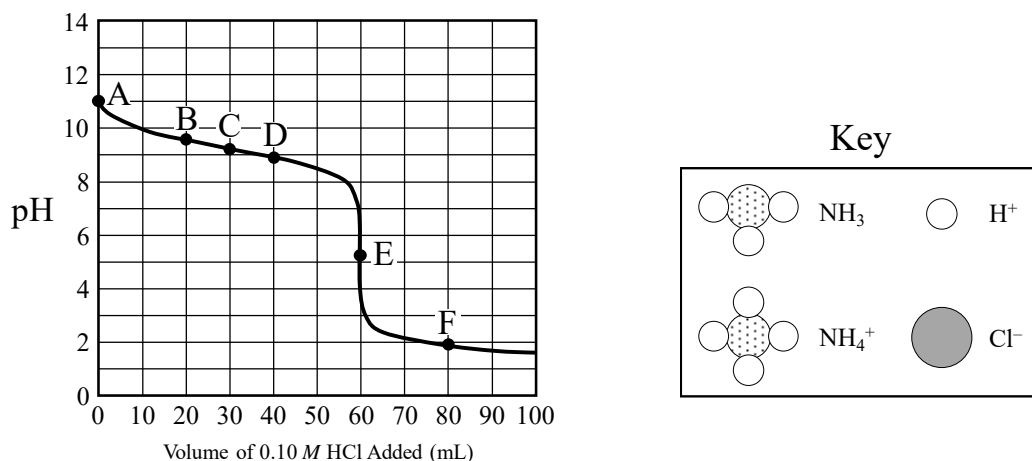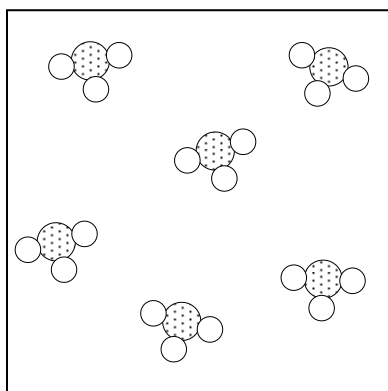
A

B

C

D

A 60.0 mL sample of 0.10 M $NH_3(aq)$ is titrated with 0.10 M $HCl(aq)$. The pH of the resulting solution is measured with a pH meter and graphed as a function of the volume of the titrant added.

pH

14
12
10
8
6
4
2
0

A
B C D
E
F

0   10   20   30   40   50   60   70   80   90   100
Volume of 0.10 M HCl Added (mL)
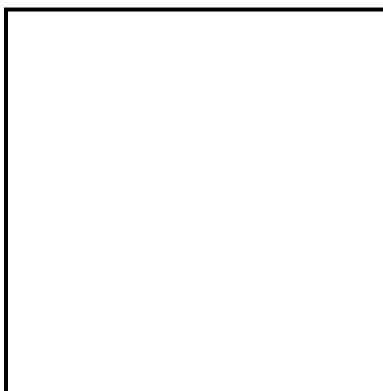
Key

NH_3 H^+

$NH_4{}^+$ Cl^-

The diagram labeled as "A" below represents a particulate representation of a small representative portion of the solution at point A of the pH curve shown above. For clarity, water molecules are not shown.

Use the symbols shown in the Key above to draw particulate representations of the solution at points B, C, D, E, and F of the pH curve shown above.
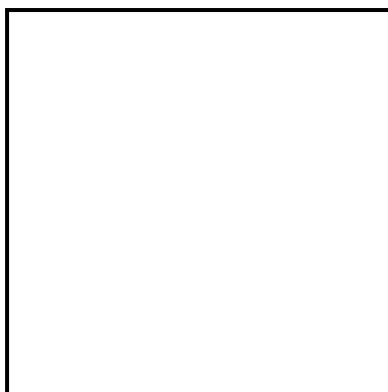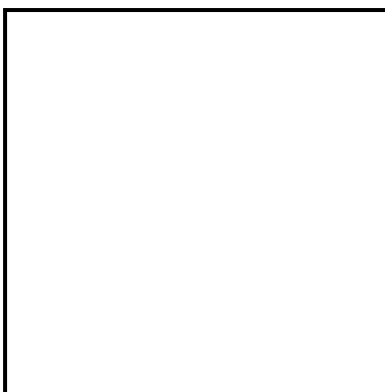
A

B

C

D

E

F