

Name:

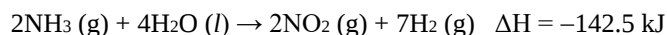
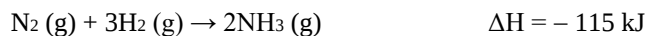
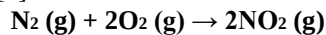
Date:

Period:

Seat #:

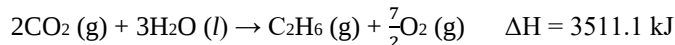
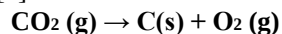
Show all work

[1] Find the ΔH for the reaction below, given the following reactions and subsequent ΔH values:



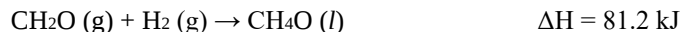
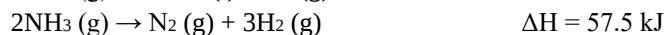
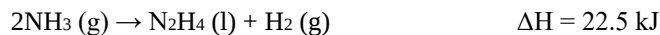
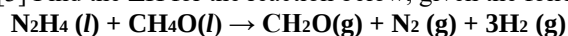
Answer = -83 kJ

[2] Find the ΔH for the reaction below, given the following reactions and subsequent ΔH values:



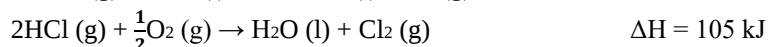
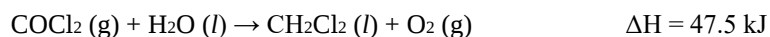
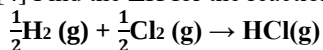
Answer = 886 kJ

[3] Find the ΔH for the reaction below, given the following reactions and subsequent ΔH values:



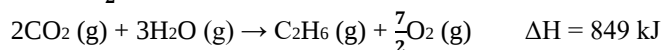
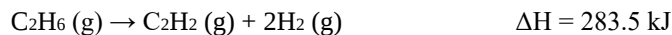
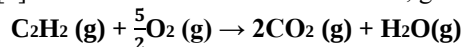
Answer = -46.2 kJ

[4] Find the ΔH for the reaction below, given the following reactions and subsequent ΔH values:



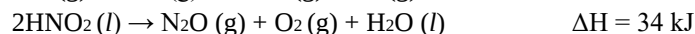
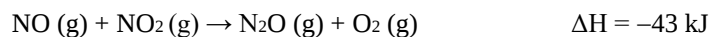
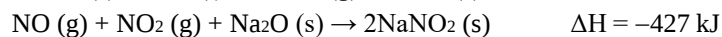
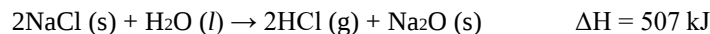
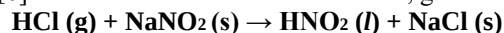
Answer = -230 kJ

[5] Find the ΔH for the reaction below, given the following reactions and subsequent ΔH values:



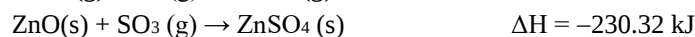
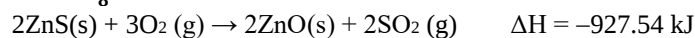
Answer = -705 kJ

[6] Find the ΔH for the reaction below, given the following reactions and subsequent ΔH values:



Answer = -78 kJ

[7] Find the ΔH for the reaction below, given the following reactions and subsequent ΔH values:



Answer = -976.03 kJ

These problems involve using heat of formation values that are found in the appendix of your textbook

[8] What is the enthalpy of the following reactions? Use equation from Equation Sheet – [Products minus reactants]

