**Name: Period: Seat#:**

**Worksheet #5**

**Directions:** Show all work in a way that would earn you credit on the AP Test! This is always the rule! Answers provided at the bottom of the last page. **SHOW WORK ON BINDER PAPER. STAPLE BINDER PAPER TO THE BACK OF THIS WORKSHEET!** Clearly label which question number your work is for.

NChO 1999

|  |  |
| --- | --- |
| **Reaction** | **ΔH** |
| Mg(s) + 2 HCl(aq) → MgCl2(aq) + H2(g) | -467 kJ mol¯1 |
| MgO(s) + 2 HCl(aq) → MgCl2(aq) + H2O(l) | -151 kJ mol¯1 |

27. According to this information, and given the fact that for water, Hf = -286 kJ mol¯1, what is Hf for MgO(s)?

 (A) -904 kJ mol¯1 (C) -334 k J mol¯1
(B) -602 kJ mol¯1 (D) -30 kJ mol¯1

NChO 1998

22. Carbon reacts with oxygen according to this equation. 2C(s) + O2(g) → 2CO(g) H = -220 kJ

 Which statements are true?

 1. The reaction is exothermic.

 2. The combustion of 0.50 mol of carbon produces 55 kJ of heat energy

 (A) 1 only (C) both 1 and 2
(B) 2 only (D) neither 1 nor 2

 24. Use these data to calculate H° for this reaction.
NO(g) + ½ O2(g) → NO2(g)

 (A) -57.0 (C) +28.5
(B) -28.5 (D) +57.0

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| Reaction  | H°, kJ mol¯1  |
| ½ N2(g) + ½ O2(g) → NO(g)  | 90.2 kJ mol¯1  |
| ½ N2(g) + O2(g) → NO2(g)  | 33.2 kJ mol¯1  |

 25. A 1.0 g sample of substance A at 100 °C is added to 100 ml of H2O at 25 °C. Using separate 100 mL portions of H2O, the procedure is repeated with subs. B and then with subs. C. How will the final temps of the water compare?

 (A) Tc > Tb > Ta (C) Ta > Tb > Tc
(B) Tb > Ta > Tc (D) Ta = Tb = Tc

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| Substance | Specific Heat |
| A | 0.60 J g-1 °C-1 |
| B | 0.40 J g-1 °C-1 |
| C | 0.20 J g-1 °C-1 |

26. How many grams of benzene, C6H6(l), must be burned in a bomb calorimeter to raise its temperature by 1.5 °C? Given: The calorimeter constant is 12.59 kJ C¯1 and the H°:combustion for C6H6 = -41.9 kJ g¯1

 (A) 0.45 g (C) 3.3 g
(B) 2.8 g (D) 8.4 g

NChO 1997

19. 30.0 mL of water at 10. °C is mixed with 50.0 mL of water at 60. °C. What is the final temp of the mixture?

 (A) 31 °C (C) 41 °C
(B) 35 °C (D) 46 °C

24. What is the value of H° for this reaction?
Fe2O3(s) + 3H2O(*l*) → 2Fe(OH)3(s)

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| --- | --- |
| Substance  | H°f  (kJ mol¯1)  |
| Fe2O3(s) | -824.2 |
| Fe(OH)3(s) | -823.0 |
| H2O(l) | -285.8 |

 (A) 35.6 kJ (C) 858.6 kJ
(B) 286 kJ (D) -536 kJ



25. When Na2S2O3 **.** 3H2O
dissolves in water, the
solution gets cold. Which
energy diagram best
represents the behavior of
this solution process?

NChO 1996

22. The standard enthalpy of formation (ΔH°f) for sodium bromide is the enthalpy change for the reaction

 (A) Na+(g) + Br¯(g) → NaBr(g)
(B) Na+(g) + Br¯(g) → NaBr(s)
(C) 2 Na(s) + Br2(g) → 2 NaBr(s)
(D) Na(s) + ½ Br2(l) → NaBr(s)

 23. Use the standard enthalpies of formation to calculate ΔH° for this reaction:
2 CrO42¯(aq) + 2 H+(aq) → Cr2O72¯(aq) + H2O(*l*)

 (A) 272.1 kJ (C) -13.7 kJ
(B) 13.7 kJ (D) -272.1 kJ

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| Substance | ΔH°f, kJ mol¯1 |
| CrO42¯(aq) | - 881.2 |
| Cr2O72¯(aq) | - 1490.3 |
| H+(aq) | 0 |
| H2O(l) | - 285.8 |

NChO 1995

21. For which of these processes is the sign of the enthalpy change different from the others?

 (A) Al2O3(s) → 2 Al(s) + 3/2 O2(g)
(B) H2O(s) → H2O(*l*)
(C) Cl2(g) → 2Cl(g)
(D) Cl(g) + e¯ → Cl¯(g)

22. The standard enthalpy of formation (H°) for nitrogen(IV) oxide is the enthalpy change for the reaction

 (A) N(g) + 2O(g) → NO2(g)
(B) ½ N2(g) + O2(g) → NO2(g)
(C) ½ N2O4(g) → NO2(g)
(D) NO(g) + ½ O2 → NO2(g)

23. In a bomb calorimeter, reactions are carried out at

 (A) constant volume. (C) 1 atm pressure and 25 °C.
(B) constant pressure. (D) 1 atm pressure and 0 °C.

26. Consider the reaction
Hg(*l*) + 2 Ag+(aq) → Hg2+(aq) + 2 Ag(s)

 What is enthalpy change for this rxn if ΔH°f for Ag+ (aq) is +105.6 kJ mol¯1 and for Hg2+ (aq) is +171.1 kJ mol¯1?

 (A) 65.5 kJ are evolved per mole of Hg.
(B) 65.5 lK are absorbed per mole of Hg.
(C) 40.1 kJ are evolved per mole of Hg.
(D) 40.1 kJ are absorbed per mole of Hg.

NChO 1994

24. A student mixes 100 mL of 0.50 M NaOH with 100 mL of 0.50 M HCl in a styrofoam cup and observes a temperature increase of T1. When she repeats this experiment using 200mL of each solution, she observes a temperature change of T2. If no heat is lost to the surroundings or absorbed by the styrofoam cup, what is true about T1 and  T2?

 (A) T2 = T1 (C) T2 = 2 T1
(B) T2 = 0.5 T1 (D) T2 = 4 T1

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| 27. Given these values of H°:  | H° |
| CS2(*l*) + 3O2(g) → CO2(g) + 2SO2(g)  | - 1077 kJ |
| H2(g) + O2(g) → H2O2(*l*)  | - 188 kJ |
| H2(g) + (1/2) O2(g) → H2O(*l*)  | - 286 kJ |

 What is the value of H° for this reaction?

CS2(*l*) + 6 H2O2(*l*) → CO2(g) + 6 H2O(*l*) + 2 SO2(g)

 (A) -1175 kJ (C) -1665 kJ
(B) -1551 kJ (D) -3921 kJ

NChO 1993

13. Which process or reaction has a positive H?

 (A) H2O(*l*) → H2O(s)
(B) 2CH3OH(*l*) + 3O2(g) → 2CO2(g) + 4H2O(*l*)
(C) CO2(s) → CO2(g)
(D) 2 Na(s) + Cl2(g) → 2 NaCl(s)

15. For the reaction H2(g) + I2(s) → 2 HI(g) Hrxn = 53.0 kJ

 What will be the value of Hrxn (in kJ) for this rxn ?

 HI(g) → ½ H2(g) + ½ I2(s)

 (A) 26.5 (C) -26.5
(B) 7.3 (D) -53.0

***Answers***

*1999*

*27) B*

*1998*

*22) C*

*24) A*

*25) C*

*26) A*

*1997*

*19) C*

*24) A*

*25) A*

*1996*

*22) D*

*23) C*

*1995*

*21) D*

*22) B*

*23) A*

*26) C*

*1994*

*24) A*

*27) C*

*1993*

*13) C*

*15) C*