Name:_____

Buffer inquiry lab

You will create a buffer in two different ways using the following chemicals.

- 0.10 M acetic acid, $HC_2H_3O_2$, Ka= 1.75×10^{-5}
- Solid sodium acetate $NaC_2H_3O_2$
- 0.10 M sodium hydroxide, NaOH

To create the buffers,

(1) Must contain 10 mL of $HC_2H_3O_2$ solution.
(2) The buffer must have a pH of 4.76

In the space provided, come up with a plan. You must include all amounts used of chemical used. When finished, come check the pH of the buffer solution with me. I will initial if you are close enough to the pH.

Buffer one	Buffer two
Containining $HC_2H_3O_2$ solution and $NaC_2H_3O_2$ solid	Containing $HC_2H_3O_2$ solution and NaOH solution
Calculations:	Calculations:
Procedure:	Procedure:
Signature	Signature