I

Name:	Date:	Period:	Seat #:
Name:	Date:	Perioa:	Seat #:

Show all work for each problem. Box your final answers.

The overall dissociation of oxalic acid, H₂C₂O₄, is represented below. The overall dissociation constant is also indicated. H₂C₂O₄ \Leftrightarrow 2 H⁺ + C₂O₄²⁻ K = 3.78 x 10⁻⁶

[1] What volume of 0 400-molar NaOH is required to neutralize completely a 5 00 x 10^{-3} mole sample of pure ovalic acid? 25.0 mJ
[1] what volume of 0.400-molar NaO11 is required to neutralize completely a 5.00 x 10 ⁻ mole sample of pure oxalic acid: 25.0 mL
121 Give the equations representing the first and second dissociations of oxalic acid
[2] Give the equations representing the mist and second dissociations of oxane deld.
$[2a]$ Calculate the value of the first dissociation constant, K_1 , for oxalic acid if the value of the second dissociation constant, K_2 , is
$6.40 \times 10^{-5}, 5.91 \times 10^{-2}$
(2) To a 0.015 metric relation of earlier and a structure of the added with the addition of C of the description of the manufactory of the structure of the
$[5]$ 10 a 0.013-molar solution of oxalic acid, a strong acid is added until the pH is 0.3. Calculate the $[C_2O_4^{-2}]$ in the resulting
solution. (Assume the change in volume is negligible.) 5.67×10^{-7} M
Al Calculate the value of the equilibrium constant K. for the reaction that occurs when solid NacCaO ₄ is discolved in water
[+] Calculate the value of the equilibrium constant, K _b , for the reaction that occurs when solid Na ₂ C ₂ O ₄ is dissolved in water.
1.56×10 ⁻⁴