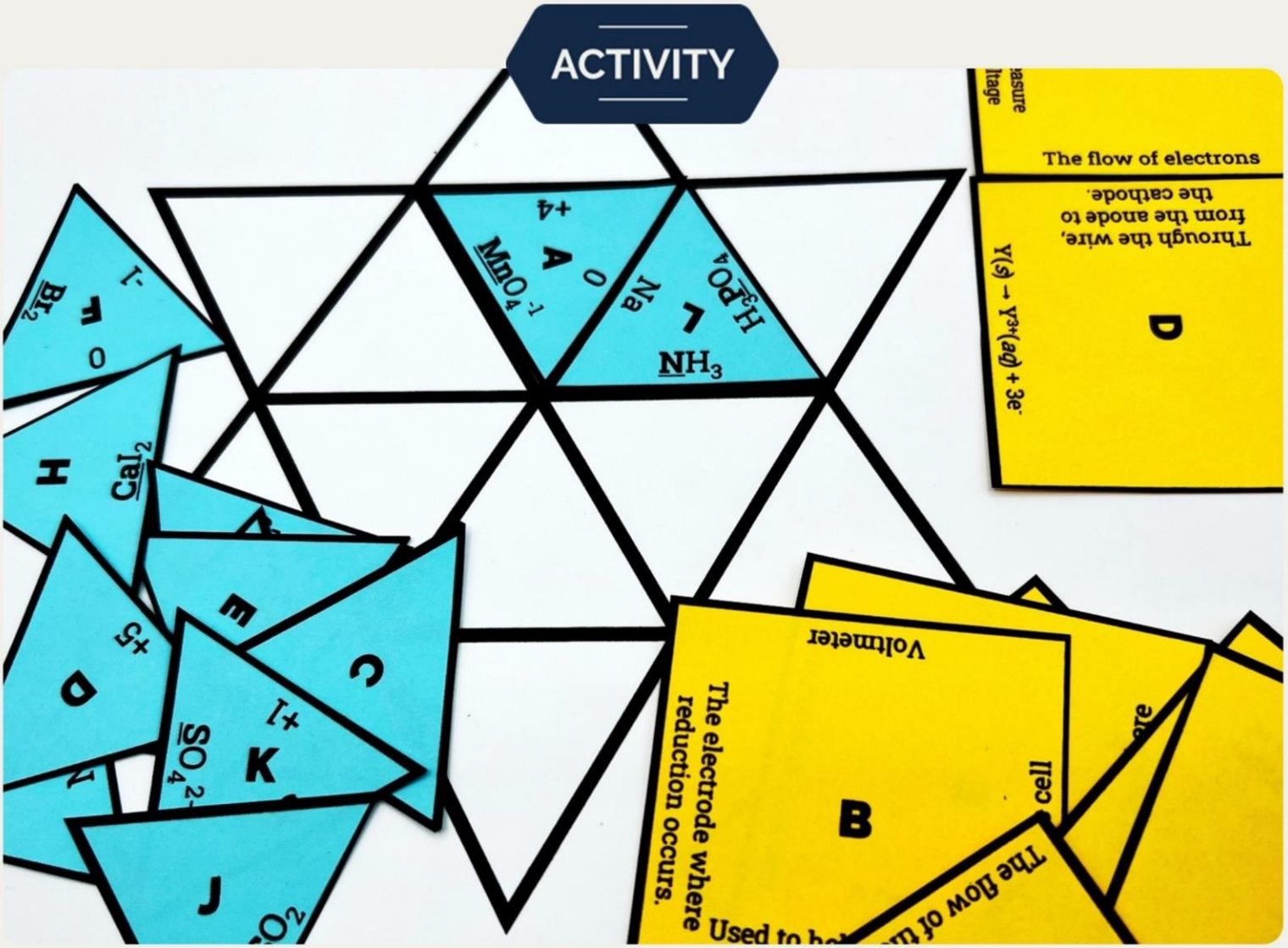
Thank you for your download!

Oxidation Numbers and Galvanic Cell Puzzles

Print and Digital

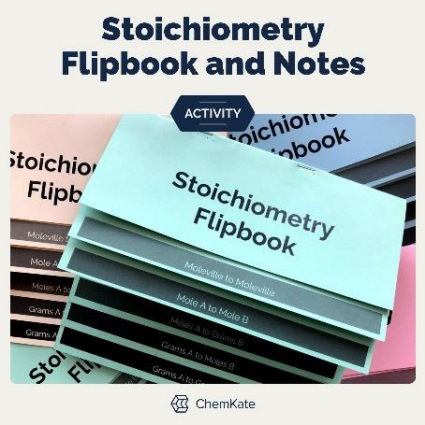


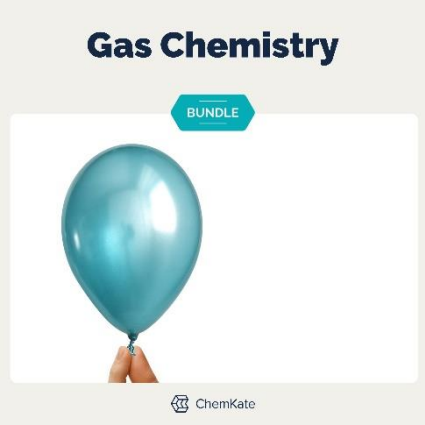
ChemKate

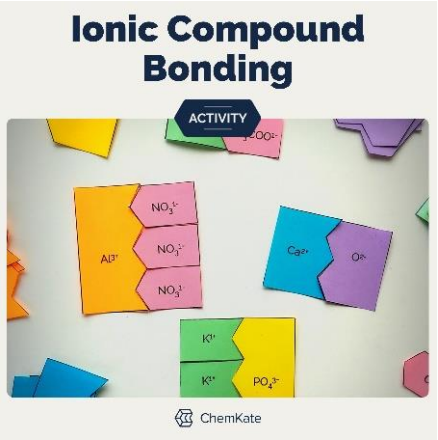
I'd love your feedback!

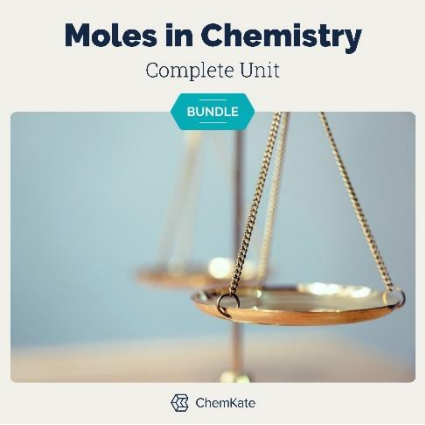
Oxidation Number and Galvanic Cell Parts Tarsia Puzzles

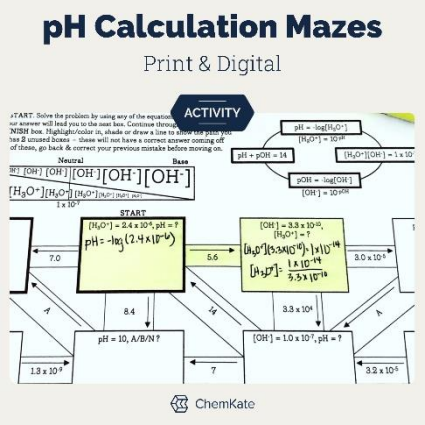
Thank you for your download! You might also be interested in the linked images below:

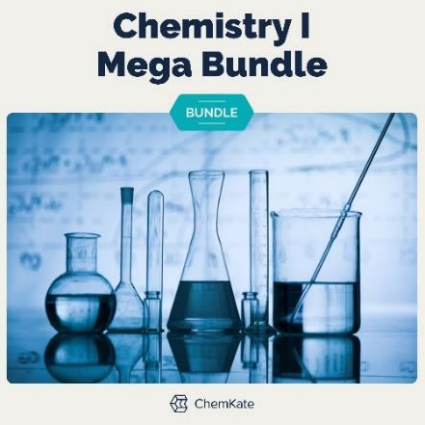












Looking for an easy, quick review of oxidation numbers and reviewing the parts and functions of a galvanic cell your redox reactions study? In low-prep print and no-prep digital, tarsia puzzles offer a great way to practice oxidation numbers and galvanic cells in a new way. Both puzzles can be used as bellringer/bellwork, a break in notes, at stations, a homework assignment, exit activity or for early finishers.

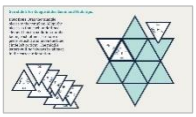
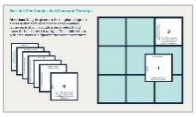
✦ This is available in my costs-savings Redox Reactions Bundle - and save *time and assurance* with all the activities such as interactive particle diagrams, labs, and *self-grading* homework found in this **Chemistry I Activity Mega Bundle** ✦

Included in these files, print and digital:

- Oxidation Number Tarsia Puzzle with optional Student Template. Topics include oxidation numbers -3 to +7.
- Galvanic Cell Tarsia Puzzle with optional Student Template. Topics include anode, cathode, salt bridge, flow of e-, flow of spectator ions in salt bridge, how concentrations change over time.
- Answer Keys

Accessing the Digital Activities

1. Be sure you are logged into the Google account you want to save these files into first. When you select the links on the next page, it will ask you to make a copy of the assignment. Select "Make a Copy".

Preview	Student File	Answer Key
	Oxidation Number Tarsia	Oxidation Number Tarsia – Answer Key
	Galvanic Cell Tarsia	Galvanic Cell Tarsia – Answer Key

2. These copies in your drive are now your Master Templates.

3. Printable PDF versions can be found after Terms of Use.

Teacher Tips

- Printable versions: Laminate and cut out the puzzle pieces for repeated use. Store them in individual baggies and all baggies in one large manila envelope or larger baggie for *easy* use in the future.

- Print out multiple copies of the answer keys for students to check their answers against can free you up to help those who need it.

- Students can work individually or together in groups of 2-3 for collaboration and test review. For digital collaboration, your students can share a Google slide deck with each other working on one document.

- Pass out the template for students to glue their tarsia pieces on to turn in or glue into their notebook.

- Do you need a clear, easy to read, updated periodic table to use while completing this activity? Periodic tables and an activity can be found in this popular Freebie: Periodic Table of Elements Color by Category.

Interested in more great resources? Click on the linked icons below:

Bd	Nm	Ml	Vs	Lb
Bonding	Naming	Mole	VSEPR	Labs

Thank you for your download!
Did you know that **you can receive credit** towards future TpT purchases by reviewing this product? I love hearing how things worked for you and I appreciate your support by leaving a review. Scroll down to the bottom of the link below:

Claim My FREE Credit!

I enjoy making resources to help you be successful in the classroom, so if you have questions or suggestions, contact me at KateCk@ChemKate.com.

Let's connect!

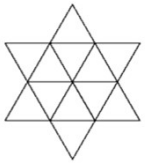
TpT Instagram Pinterest

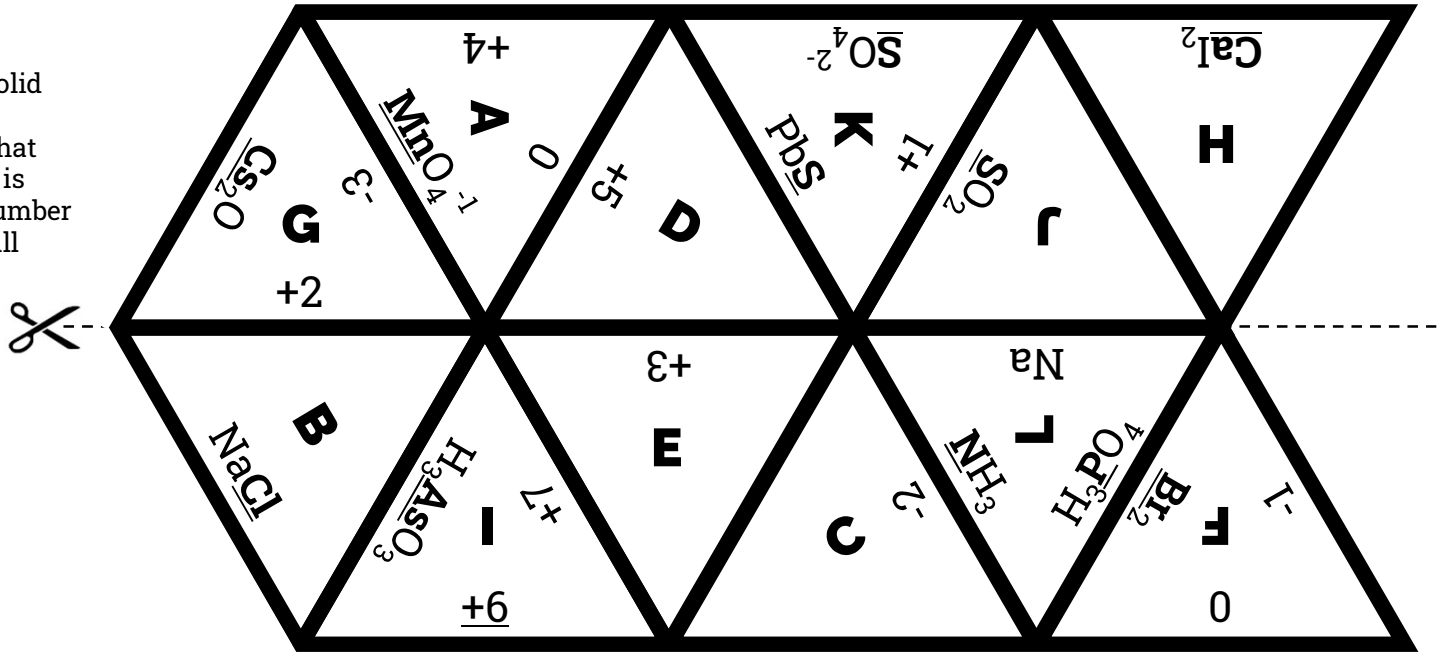
Terms of Use:

Thank you!

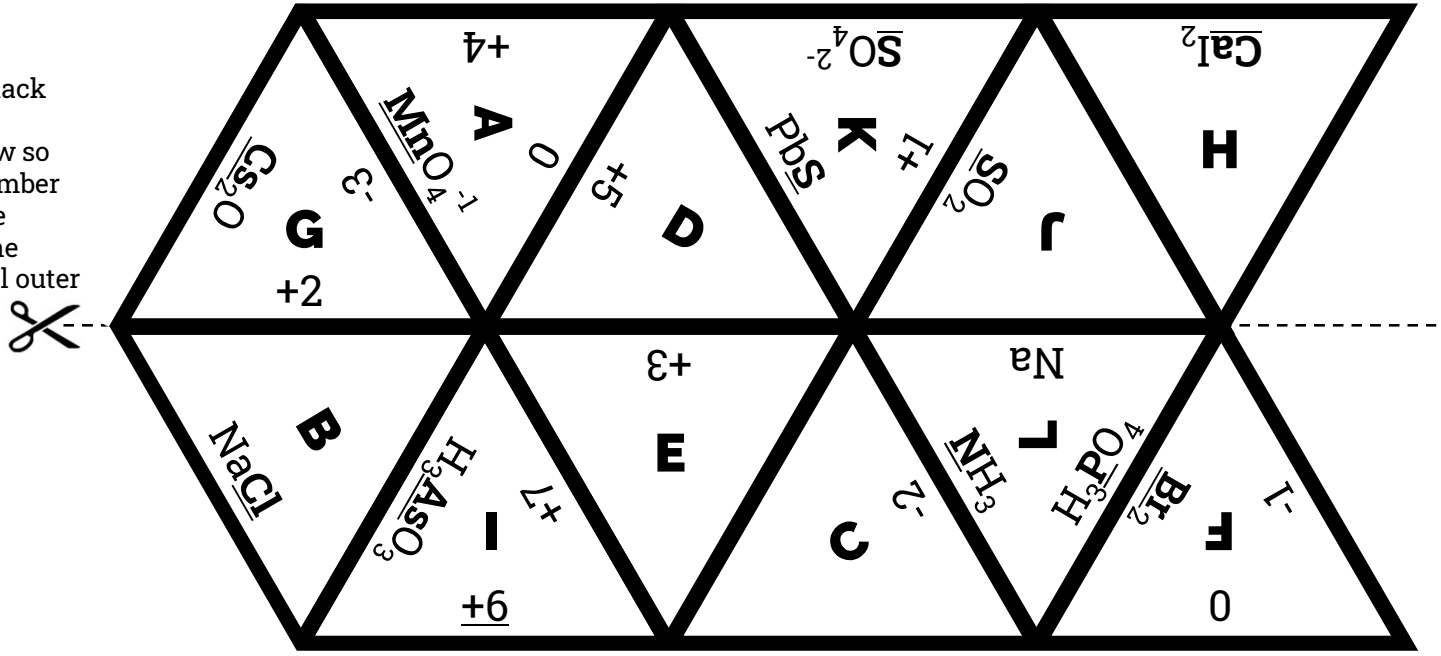
Oxidation Number Tarsia Puzzle

Directions:
1. Cut along all of the black solid lines.
2. Build the shape below so that each oxidation number edge is facing the same oxidation number as the underlined element. All outer edges will be blank.



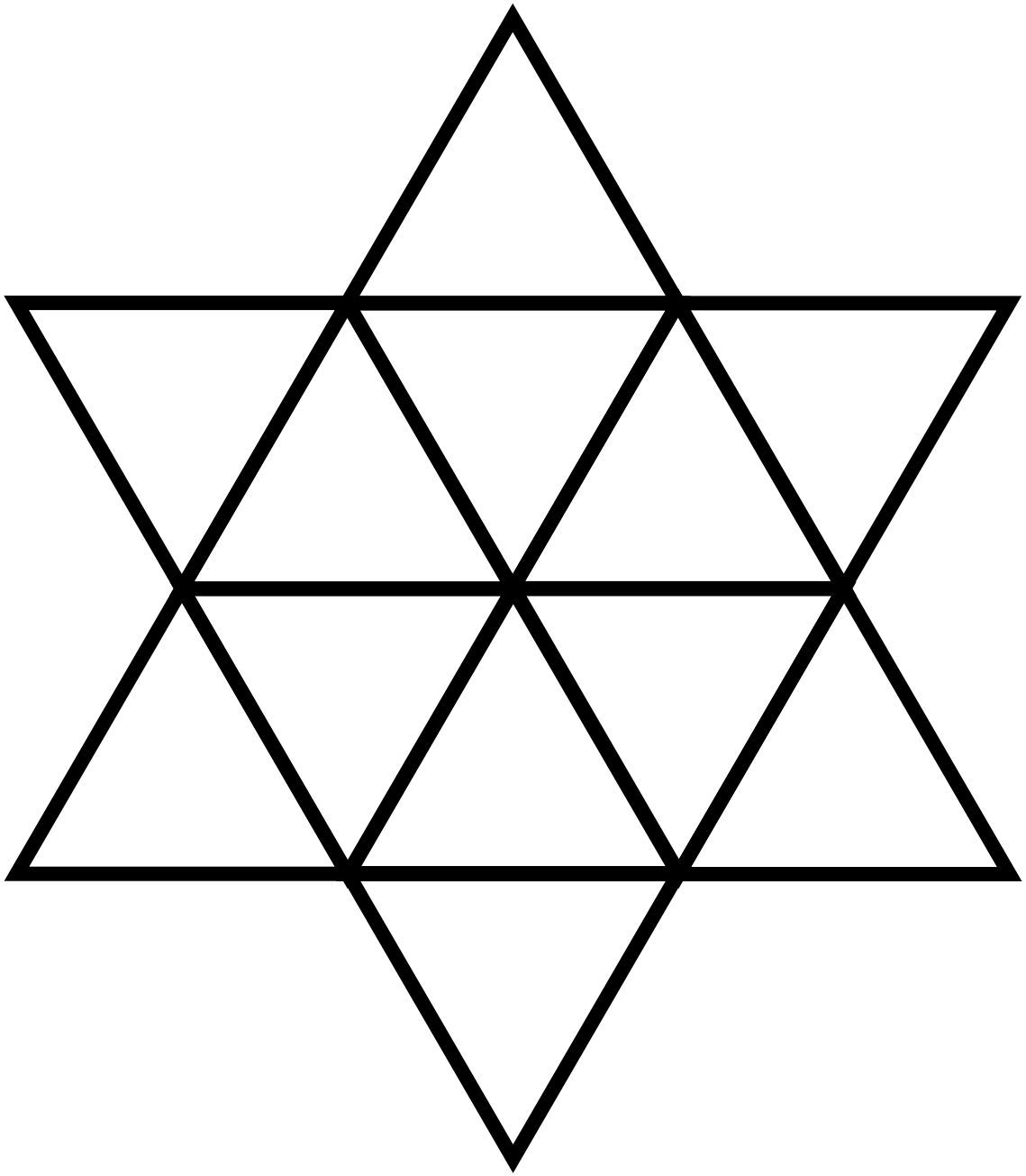


Directions:
1. Cut along all of the black solid lines.
2. Build the shape below so that each oxidation number edge is facing the same oxidation number as the underlined element. All outer edges will be blank.



Oxidation Numbers Tarsia Template

Name: _____



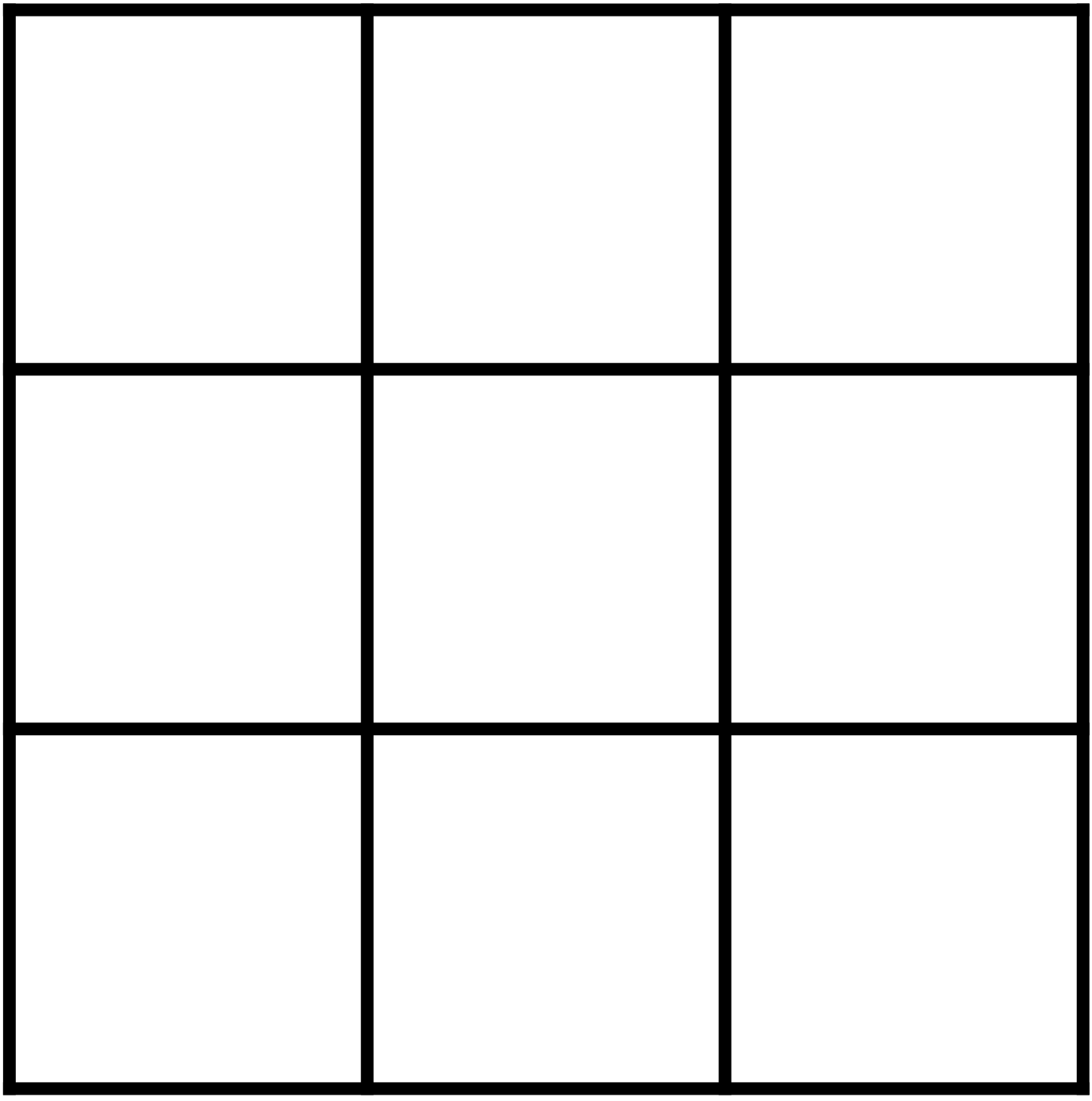
Galvanic Cell Square Tarsia Pieces To Cut

Directions:
1. Cut along all of the black solid lines.
2. Build the shape below so that each edge has the same answers facing each other. All outer edges will be blank.

Oxidation half-reaction **E** A redox reaction that is used to generate electricity. The electrode where oxidation occurs.	$Y(s) \rightarrow Y^{3+}(aq) + 3e^-$ **D** Through the wire, from the anode to the cathode.	The flow of electrons **H** $X^{2+}(aq) + 2e^- \rightarrow X(s)$ Used to measure a cell's voltage
Cathode **I** Through the salt bridge, toward the cathode. Anode solution	Anode **A** This solution decreases in the concentration of the metal ion as the galvanic cell operates.	Galvanic cell **B** The electrode where reduction occurs. Used to help maintain electrical neutrality by use of mobile ions Voltmeter
Through the salt bridge, toward the anode. **C** This solution increases in the concentration of the metal ion as the galvanic cell operates.	Reduction half-reaction **G** The flow of the spectator cation	Cathode solution **F** The flow of the spectator anion Salt bridge

Galvanic Cell Square Tarsia Template Name: _____ Pd: ____

Directions:
1. Use the template below to help you build your square tarsia.
2. Build the shape below so that each edge has the same answers facing each other. All outer edges will be blank.



Oxidation Numbers Tarsia Puzzle Answer Key

Galvanic Cell Square Tarsia Answer Key

A This solution decreases in the concentration of the metal ion as the galvanic cell operates. Anode	**F** The flow of the spectator anion Salt bridge	**C** This solution increases in the concentration of the metal ion as the galvanic cell operates. Through the salt bridge, toward the anode.
E A redox reaction that is used to generate electricity. Oxidation half-reaction	**B** The electrode where reduction occurs. Galvanic cell	**I** Anode solution Cathode Through the salt bridge, toward the cathode.
D Through the wire, from the anode to the cathode.	**H** $X^{2+}(aq) + 2e^- \rightarrow X(s)$ The flow of electrons	**G** The flow of the spectator cation Reduction half-reaction

The electrode where oxidation occurs. $Y(s) \rightarrow Y^{3+}(aq) + 3e^-$	Used to help maintain electrical neutrality by use of mobile ions Used to measure a cell's voltage	Cathode solution Voltmeter