

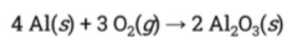
# Redox Reactions

Print, Digital, and Editable

## ACTIVITY

Question 1

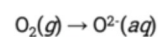
1) What is oxidized in the reaction below? \*



- $\text{O}^{2-}$
- $\text{Al}^{3+}$
- $\text{O}_2$
- $\text{Al}$

Question 9

9) Which of the following shows the balanced half-reaction of the one below? \*



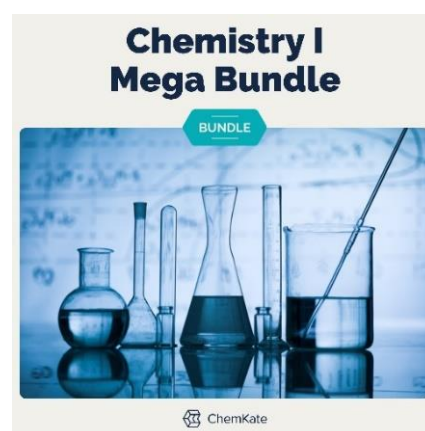
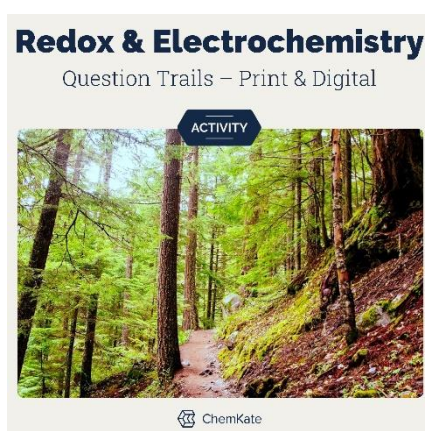
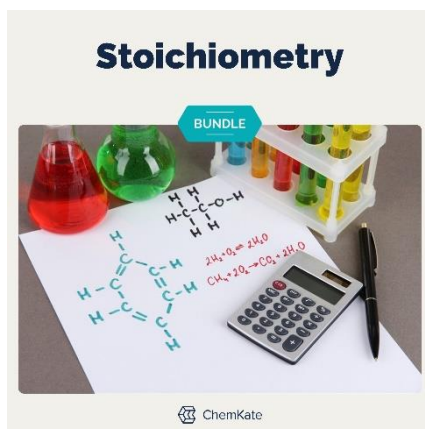
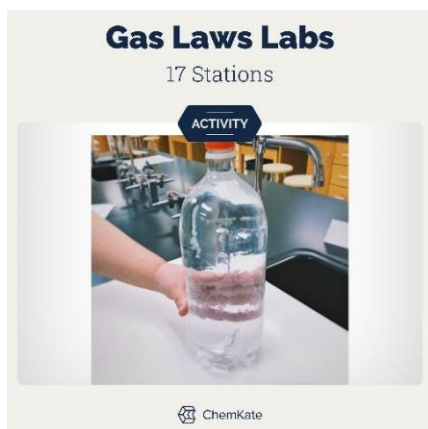
- $2e^- + \text{O}_2 \rightarrow \text{O}^{2-}$
- $\text{O}_2 \rightarrow 2 \text{O}^{2-} + 4e^-$
- $4e^- + \text{O}_2 \rightarrow 2 \text{O}^{2-}$
- $\text{O}_2 \rightarrow \text{O}^{2-} + 2e^-$



ChemKate

# Redox Reactions Image Reveal – Print and Digital

Thank you for your download! You might also be interested in the [linked](#) images below:



This *no-prep, editable* redox reactions practice, *in print and self-grading digital format*, engages students solving for what is oxidized, reduced, and balancing half-reactions. Students can complete the redox problems with a Google Form that *gives students instant feedback* and reveals a pixelated image as they work. Great for entrance/exit tickets, homework, assessments, or for early finishers.

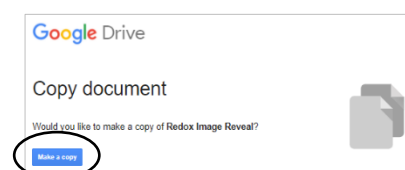
◆ This is available in my costs-savings [Redox Reactions Bundle](#), the [Chemistry Image Reveal Bundle](#) and save *time and assurance* with all the activities found in this [Chemistry I MEGA Activity Bundle](#) - Visit this bundle to see a wide variety of guided inquiry activities, graphic organizers and digital practice and application ◆

Topics:



- Identifying what is oxidized or reduced in full, half-reactions, and net ionic equations
- Balancing half-reactions
- Identifying a half-reaction as oxidation or reduction

## Accessing the Digital Activities

1. Be sure you are logged into the Google account you want to save these files into first. When you select the link(s) *on the next page*, it will ask you to make a copy of the assignment. Select "Make a copy".



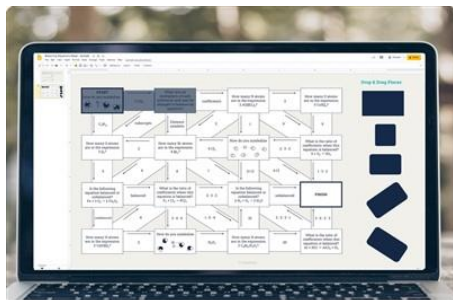
## Redox Reactions Image Reveal cont.

Preview	Digital File Link(s)	Background Info
	<a href="#">Redox Reactions Image Reveal Google Form</a>	Current editable Form settings ⚙️ <ul style="list-style-type: none"><li>• General &gt; Limited to 1 response</li><li>• Quizzes &gt; Release grade &gt; Immediately after each submission and show Missed questions</li><li>• For more help from Google with how to edit google forms for your specific needs, see tips from Google™ <a href="#">here</a>.</li></ul>
	<a href="#">Redox Reactions Editable Worksheet</a>	These slides can be downloaded as a PowerPoint. Once copied, select File > Download > Microsoft PowerPoint (.pptx)

2. These copies in your drive are now your Master Templates. I would recommend changing the name of the file and organizing the file into a folder so that you can easily access it later.

### Receive this Freebie!

Get this exclusive *no-prep* balancing equation maze - *in print and digital* formats that are *self-checking and easy to grade*, as well as receiving tips, ideas, and resources periodically sent to you.



[Get My Maze!](#)

Interested in more great resources? Click on the linked icons below:

**Bd**

Bonding

**Nm**

Naming

**Ml**

Mole

**Vs**

VSEPR

**Lb**

Labs

## Redox Reactions Image Reveal cont.

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Photography by [8926](#)  
All graphics by © ChemKate

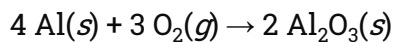
Thank you!



Directions:

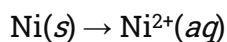
Complete each below.

1) What is oxidized in the reaction below?



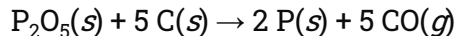
\_\_\_\_\_ is oxidized

2) Does this unbalanced half-reaction show oxidation or reduction?



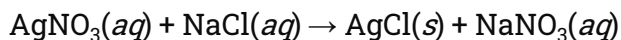
\_\_\_\_\_

3) What is reduced in the reaction below?



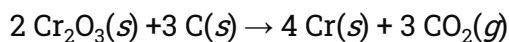
\_\_\_\_\_ is reduced

4) Is the reaction below a redox reaction?



\_\_\_\_\_

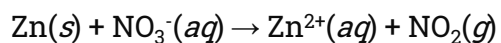
5) What is oxidized and what is reduced in the equation below?



\_\_\_\_\_ is oxidized

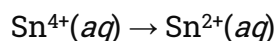
\_\_\_\_\_ is reduced

6) What is oxidized in the reaction below?



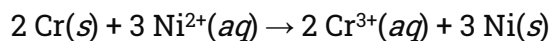
\_\_\_\_\_ is oxidized

7) How many electrons and on which side of the reaction are needed to balance this half-reaction?



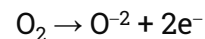
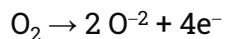
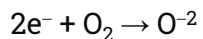
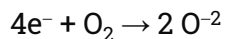
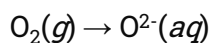
\_\_\_\_\_

8) What is reduced in the reaction below?

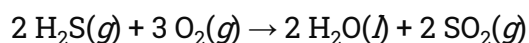


\_\_\_\_\_ is reduced

9) Circle the equation that shows the balanced half-reaction of the one below:



10) What is neither oxidized nor reduced in the reaction below?

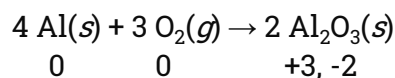


\_\_\_\_\_ is neither oxidized or reduced

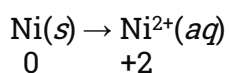
Directions:

Complete each below.

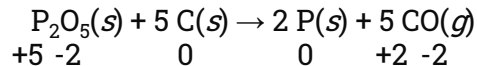
1) What is oxidized in the reaction below?

Al is oxidized

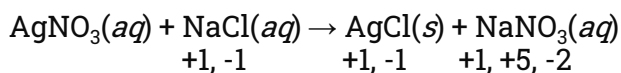
2) Does this unbalanced half-reaction show oxidation or reduction?

oxidation

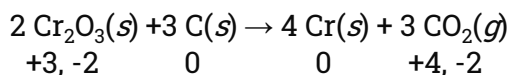
3) What is reduced in the reaction below?

P<sup>+5</sup> is reduced

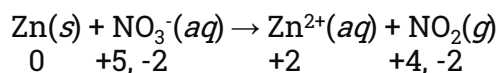
4) Is the reaction below a redox reaction?

No

5) What is oxidized and what is reduced in the equation below?

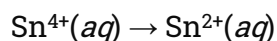
C is oxidizedCr<sup>+3</sup> is reduced

6) What is oxidized in the reaction below?



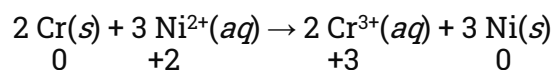
Zn is oxidized

7) How many electrons and on which side of the reaction are needed to balance this half-reaction?



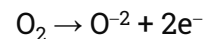
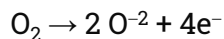
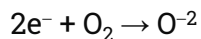
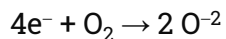
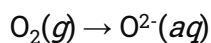
2e<sup>-</sup> added to the reactant side.

8) What is reduced in the reaction below?

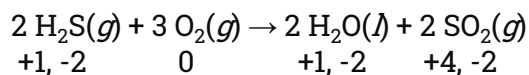


Ni<sup>2+</sup> is reduced

9) Circle the equation that shows the balanced half-reaction of the one below:



10) What is neither oxidized nor reduced in the reaction below?



H<sup>+</sup> is neither oxidized or reduced