N6 – Entropy

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| **Key Concepts About Entropy** |
| **#1 -** a thermodynamic function that increases as the number of energetically equivalent ways of arranging the components increases, S. |
| **#2 –** Random systems have more energy dispersal and are more energetically stable, lower energy, than ordered systems. Therefore, entropy change is favorable when the result is more energy dispersal, when there are more microstate arrangements. When ∆S° is positive.  |
| **#3 –** Increase in entropy of the universe is the driving force for spontaneous reactions.  |
| **#4 –** Nature proceeds toward the states that have the highest probabilities of existing. |

**Practice Problem:**For the Rx: **2NO(g) + O2(g) 🡪 2NO2(g)**

ΔS°rxn= -146.5 J/mol•K

Calculate the standard molar entropy of O2(g).

ΔS°NO(g) = 210.8 J/mol•K , ΔS°NO2(g) = 240.1 J/mol•K

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| **Some Changes that Increase Entropy** |
| **#1 -** Products are in a more dispersed  arrangement. |
| **#2 -** Larger numbers of product  molecules than reactant molecules. |
| **#3 -** Rxn’s that have an increase in  temperature (exothermic). |
| **#4 –** Products that have more  degrees of movement. |
| **#5 –** Products that have more  molecular complexity. |

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