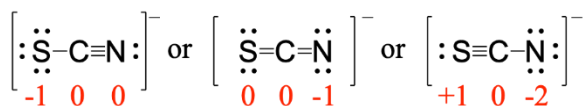
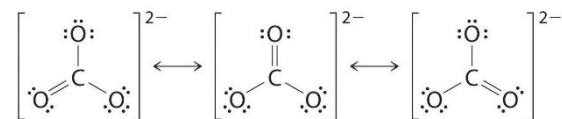
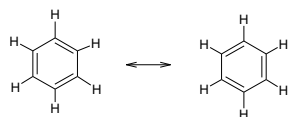


N21 – Formal Charge and VSEPR

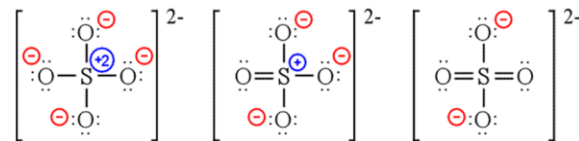
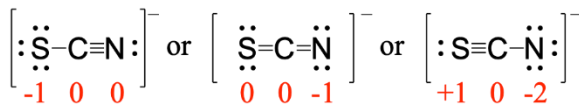
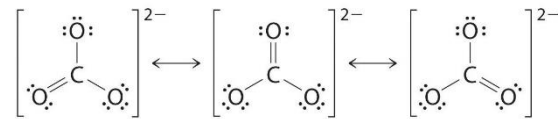
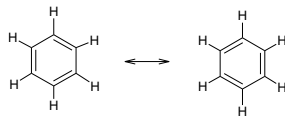


X + E	Overall Structure (Electronic Geometry)	Forms
2	Linear	AX ₂
3	Trigonal Planar	AX ₃ , AX ₂ E
4	Tetrahedral	AX ₄ , AX ₃ E, AX ₂ E ₂
5	Trigonal bipyramidal	AX ₅ , AX ₄ E, AX ₃ E ₂ , AX ₂ E ₃
6	Octahedral	AX ₆ , AX ₅ E, AX ₄ E ₂

“Rules” for Drawing Lewis Structures

- 1) **Count** and sum valence electrons.
- 2) **Place** your atoms.
- 3) Bond all atoms w/ a **single bond** (try simplest way 1st).
- 4) Give all atoms a **full shell**.
- 5) **Re-count** the electrons you used.
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 - Used **too few**? Put extras on the central atom.
 - Used **too many**? Try double or triple bonds to fix.

N21 – Formal Charge and VSEPR

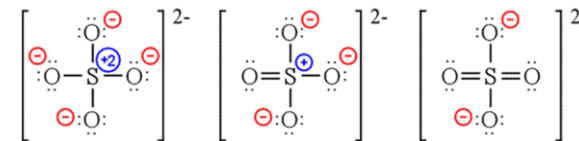
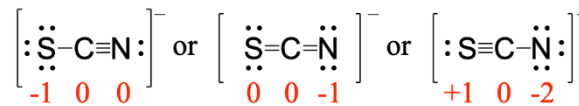
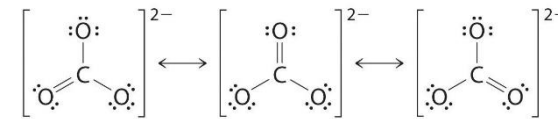
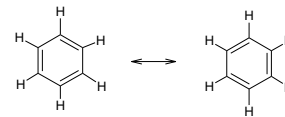


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