**Dougherty Valley HS AP Chemistry**

**WORKSHEET #1**

**Bonding – Lewis Structures (molecular)**

**Name: Date: Period: Seat #:**

Indicate the # of **VALENCE** electrons for each species. Write the correct Lewis electron-dot structure for each. Note the shape of the molecule (for compounds only)

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| **F**# of valence e–‘s = \_\_\_\_ | **O**# of valence e–‘s = \_\_\_\_ | **K**# of valence e–‘s = \_\_\_\_ | **Al**# of valence e–‘s = \_\_\_\_ |
| **F—**# of valence e–‘s = \_\_\_\_ | **O2—**# of valence e–‘s = \_\_\_\_ | **K+**# of valence e–‘s = \_\_\_\_ | **Al3+**# of valence e–‘s = \_\_\_\_ |
| **F2**# of valence e–‘s = \_\_\_\_ | **H2**# of valence e–‘s = \_\_\_\_ | **HF**# of valence e–‘s = \_\_\_\_ | **NH3**# of valence e–‘s = \_\_\_\_ |
| **CH4**# of valence e–‘s = \_\_\_\_ | **NF3**# of valence e–‘s = \_\_\_\_ | **SiF4**# of valence e–‘s = \_\_\_\_ | **C2H6**# of valence e–‘s = \_\_\_\_ |
| **MgH2**# of valence e–‘s = \_\_\_\_ | **LiH**# of valence e–‘s = \_\_\_\_ | **AlH3**# of valence e–‘s = \_\_\_\_ | **BH3**# of valence e–‘s = \_\_\_\_ |

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| **C2H4**# of valence e–‘s = \_\_\_\_ | **C2F4**# of valence e–‘s = \_\_\_\_ | **CO**# of valence e–‘s = \_\_\_\_ | **O2**# of valence e–‘s = \_\_\_\_ |
| **CO2**# of valence e–‘s = \_\_\_\_ | **C2H2**# of valence e–‘s = \_\_\_\_ | **N2**# of valence e–‘s = \_\_\_\_ | **HCN**# of valence e–‘s = \_\_\_\_ |
| **CN—**# of valence e–‘s = \_\_\_\_ | **SO42—**# of valence e–‘s = \_\_\_\_ | **PO43—**# of valence e–‘s = \_\_\_\_ | **ClO3—**# of valence e–‘s = \_\_\_\_ |
| **CO32—**# of valence e–‘s = \_\_\_\_ | **NO3—**# of valence e–‘s = \_\_\_\_ | **SO2**# of valence e–‘s = \_\_\_\_ | **O3**# of valence e–‘s = \_\_\_\_ |
| **SF6**# of valence e–‘s = \_\_\_\_ | **XeF4**# of valence e–‘s = \_\_\_\_ | **PCl5**# of valence e–‘s = \_\_\_\_ | **SeF4**# of valence e–‘s = \_\_\_\_ |