**Dougherty Valley HS AP Chemistry**

**WORKSHEET #1**

**Bonding – Lewis Structures (molecular)**

**Name: Date: Period: Seat #:**

Indicate the # of **VALENCE** electrons for each species. Write the correct Lewis electron-dot structure for each. Note the shape of the molecule (for compounds only)

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| **F**  # of valence e–‘s = \_\_\_\_ | **O**  # of valence e–‘s = \_\_\_\_ | **K**  # of valence e–‘s = \_\_\_\_ | **Al**  # of valence e–‘s = \_\_\_\_ |
| **F—**  # of valence e–‘s = \_\_\_\_ | **O2—**  # of valence e–‘s = \_\_\_\_ | **K+**  # of valence e–‘s = \_\_\_\_ | **Al3+**  # of valence e–‘s = \_\_\_\_ |
| **F2**  # of valence e–‘s = \_\_\_\_ | **H2**  # of valence e–‘s = \_\_\_\_ | **HF**  # of valence e–‘s = \_\_\_\_ | **NH3**  # of valence e–‘s = \_\_\_\_ |
| **CH4**  # of valence e–‘s = \_\_\_\_ | **NF3**  # of valence e–‘s = \_\_\_\_ | **SiF4**  # of valence e–‘s = \_\_\_\_ | **C2H6**  # of valence e–‘s = \_\_\_\_ |
| **MgH2**  # of valence e–‘s = \_\_\_\_ | **LiH**  # of valence e–‘s = \_\_\_\_ | **AlH3**  # of valence e–‘s = \_\_\_\_ | **BH3**  # of valence e–‘s = \_\_\_\_ |

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| **C2H4**  # of valence e–‘s = \_\_\_\_ | **C2F4**  # of valence e–‘s = \_\_\_\_ | **CO**  # of valence e–‘s = \_\_\_\_ | **O2**  # of valence e–‘s = \_\_\_\_ |
| **CO2**  # of valence e–‘s = \_\_\_\_ | **C2H2**  # of valence e–‘s = \_\_\_\_ | **N2**  # of valence e–‘s = \_\_\_\_ | **HCN**  # of valence e–‘s = \_\_\_\_ |
| **CN—**  # of valence e–‘s = \_\_\_\_ | **SO42—**  # of valence e–‘s = \_\_\_\_ | **PO43—**  # of valence e–‘s = \_\_\_\_ | **ClO3—**  # of valence e–‘s = \_\_\_\_ |
| **CO32—**  # of valence e–‘s = \_\_\_\_ | **NO3—**  # of valence e–‘s = \_\_\_\_ | **SO2**  # of valence e–‘s = \_\_\_\_ | **O3**  # of valence e–‘s = \_\_\_\_ |
| **SF6**  # of valence e–‘s = \_\_\_\_ | **XeF4**  # of valence e–‘s = \_\_\_\_ | **PCl5**  # of valence e–‘s = \_\_\_\_ | **SeF4**  # of valence e–‘s = \_\_\_\_ |