Dougherty Valley HS AP Chemistry
Bonding – Lewis Structures (molecular)

WORKSHEET #1

Name: **Date:** **Period:** **Seat #:**

Indicate the # of **VALENCE** electrons for each species. Write the correct Lewis electron-dot structure for each. Note the shape of the molecule (for compounds only)

F # of valence e⁻'s = ____	**O** # of valence e⁻'s = ____	**K** # of valence e⁻'s = ____	**Al** # of valence e⁻'s = ____
F^- # of valence e⁻'s = ____	**O^{2-}** # of valence e⁻'s = ____	**K^+** # of valence e⁻'s = ____	**Al^{3+}** # of valence e⁻'s = ____
F_2 # of valence e⁻'s = ____	**H_2** # of valence e⁻'s = ____	**HF** # of valence e⁻'s = ____	**NH_3** # of valence e⁻'s = ____
CH_4 # of valence e⁻'s = ____	**NF_3** # of valence e⁻'s = ____	**SiF_4** # of valence e⁻'s = ____	**C_2H_6** # of valence e⁻'s = ____
MgH_2 # of valence e⁻'s = ____	**LiH** # of valence e⁻'s = ____	**AlH_3** # of valence e⁻'s = ____	**BH_3** # of valence e⁻'s = ____

C_2H_4 # of valence e^-'s = ____	**C_2F_4** # of valence e^-'s = ____	**CO** # of valence e^-'s = ____	**O_2** # of valence e^-'s = ____
CO_2 # of valence e^-'s = ____	**C_2H_2** # of valence e^-'s = ____	**N_2** # of valence e^-'s = ____	**HCN** # of valence e^-'s = ____
CN^- # of valence e^-'s = ____	**SO_4^{2-}** # of valence e^-'s = ____	**PO_4^{3-}** # of valence e^-'s = ____	**ClO_3^-** # of valence e^-'s = ____
CO_3^{2-} # of valence e^-'s = ____	**NO_3^-** # of valence e^-'s = ____	**SO_2** # of valence e^-'s = ____	**O_3** # of valence e^-'s = ____
SF_6 # of valence e^-'s = ____	**XeF_4** # of valence e^-'s = ____	**PCl_5** # of valence e^-'s = ____	**SeF_4** # of valence e^-'s = ____