**Dougherty Valley HS AP Chemistry**

**WORKSHEET #2**

**Bonding – AXE, Lewis, and more (molecular)**

**Name: Date: Period: Seat #:**

Draw and State the following each compound below:

1. AXE formula
2. Molecular Geometry
3. Polarity
4. Total # of valence electrons

|  |  |  |  |
| --- | --- | --- | --- |
| **AlCl3**  # of valence e–‘s = \_\_\_\_ | **BCl3**  # of valence e–‘s = \_\_\_\_ | **XeO4**  # of valence e–‘s = \_\_\_\_ | **NO2**  # of valence e–‘s = \_\_\_\_ |
| **NO2+**  # of valence e–‘s = \_\_\_\_ | **PCl3**  # of valence e–‘s = \_\_\_\_ | **ClO2—**  # of valence e–‘s = \_\_\_\_ | **CCl4**  # of valence e–‘s = \_\_\_\_ |
| **XeF4**  # of valence e–‘s = \_\_\_\_ | **ClO4—**  # of valence e–‘s = \_\_\_\_ | **PCl5**  # of valence e–‘s = \_\_\_\_ | **O3**  # of valence e–‘s = \_\_\_\_ |
| **SCl2**  # of valence e–‘s = \_\_\_\_ | **SF4**  # of valence e–‘s = \_\_\_\_ | **IF4—**  # of valence e–‘s = \_\_\_\_ | **SiCl4**  # of valence e–‘s = \_\_\_\_ |
| **GaH3**  # of valence e–‘s = \_\_\_\_ | **SF6**  # of valence e–‘s = \_\_\_\_ | **OCS**  # of valence e–‘s = \_\_\_\_ | **ClF2+**  # of valence e–‘s = \_\_\_\_ |