**Dougherty Valley HS AP Chemistry**

**WORKSHEET #2**

**Bonding – AXE, Lewis, and more (molecular)**

**Name: Date: Period: Seat #:**

Draw and State the following each compound below:

1. AXE formula
2. Molecular Geometry
3. Polarity
4. Total # of valence electrons

|  |  |  |  |
| --- | --- | --- | --- |
| **AlCl3**# of valence e–‘s = \_\_\_\_ | **BCl3**# of valence e–‘s = \_\_\_\_ | **XeO4**# of valence e–‘s = \_\_\_\_ | **NO2**# of valence e–‘s = \_\_\_\_ |
| **NO2+**# of valence e–‘s = \_\_\_\_ | **PCl3**# of valence e–‘s = \_\_\_\_ | **ClO2—**# of valence e–‘s = \_\_\_\_ | **CCl4**# of valence e–‘s = \_\_\_\_ |
| **XeF4**# of valence e–‘s = \_\_\_\_ | **ClO4—**# of valence e–‘s = \_\_\_\_ | **PCl5**# of valence e–‘s = \_\_\_\_ | **O3**# of valence e–‘s = \_\_\_\_ |
| **SCl2**# of valence e–‘s = \_\_\_\_ | **SF4**# of valence e–‘s = \_\_\_\_ | **IF4—**# of valence e–‘s = \_\_\_\_ | **SiCl4**# of valence e–‘s = \_\_\_\_ |
| **GaH3**# of valence e–‘s = \_\_\_\_ | **SF6**# of valence e–‘s = \_\_\_\_ | **OCS**# of valence e–‘s = \_\_\_\_ | **ClF2+**# of valence e–‘s = \_\_\_\_ |