Name: **Date:**   **Period:**   **Seat #:**

Draw and State the following each compound below:
   a. AXE formula
   b. Molecular Geometry
   c. Polarity
   d. Total # of valence electrons

AlCl₃ # of valence e⁻'s = _____	**BCl₃** # of valence e⁻'s = _____	**XeO₄** # of valence e⁻'s = _____	**NO₂** # of valence e⁻'s = _____
NO₂⁺ # of valence e⁻'s = _____	**PCl₃** # of valence e⁻'s = _____	**ClO₂⁻** # of valence e⁻'s = _____	**CCl₄** # of valence e⁻'s = _____
XeF₄ # of valence e⁻'s = _____	**ClO₄⁻** # of valence e⁻'s = _____	**PCl₅** # of valence e⁻'s = _____	**O₃** # of valence e⁻'s = _____
SCl₂ # of valence e⁻'s = _____	**SF₄** # of valence e⁻'s = _____	**IF₄⁻** # of valence e⁻'s = _____	**SiCl₄** # of valence e⁻'s = _____
GaH₃ # of valence e⁻'s = _____	**SF₆** # of valence e⁻'s = _____	**OCS** # of valence e⁻'s = _____	**ClF₂⁺** # of valence e⁻'s = _____