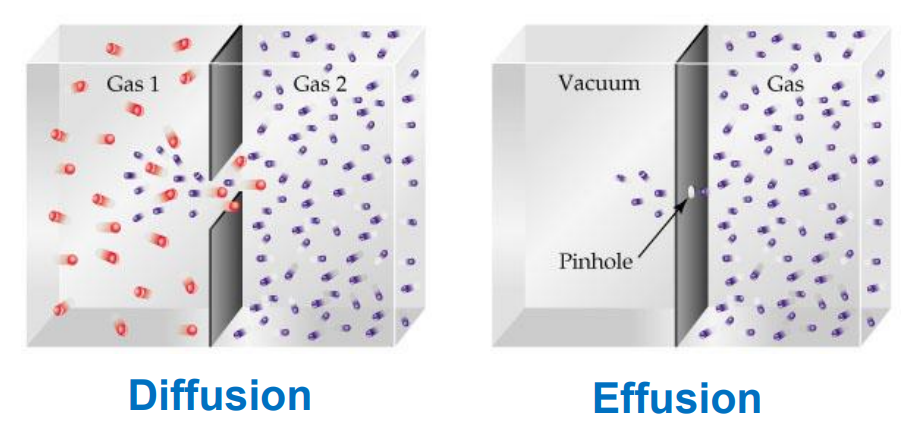
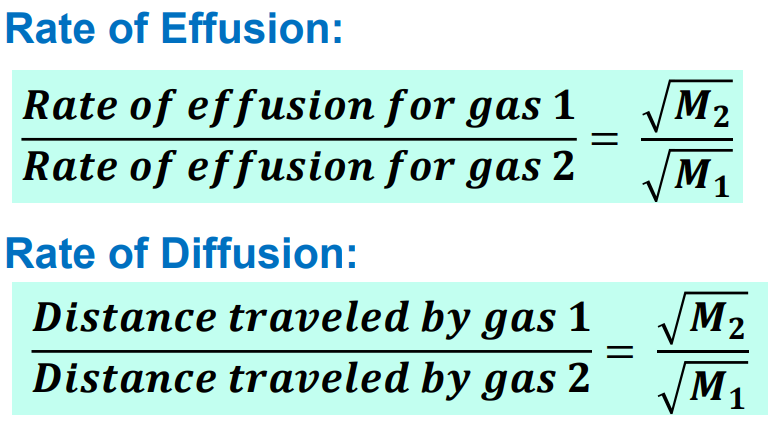
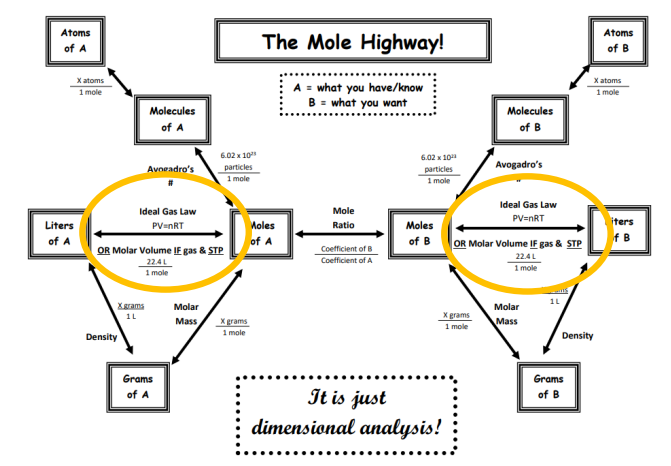
N27 – Gas Density and More

N26 – Ideal Gases and Laws

**Ideal Gas Law**

**Molar Mass Kitty**

**Boyle’s Law**

Chart, line chart

Description automatically generated

**Charles’s Law**

Chart, line chart

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**Gay-Lussac’s Law**

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**Avogadro’s Law**

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**Dalton’s Partial Pressure**

**Mole Fraction**

N25 – Gases Review

**KMT = Kinetic Molecular Theory**

1. Gases consist of large numbers of tiny particles that are far apart relative to their size   
   – the volume of each gas molecule is considered negligible, they are treated as point particles.
2. Gas particles undergo elastic collisions   
   – meaning they do not lose energy when colliding.
3. Gas particles are in a constant, rapid, straight line, “chaotic” motion   
   – they possess kinetic energy (motion energy).
4. Gases are “Ideal Gases” – meaning they do not interact with each other.   
   – There are no forces of attraction or repulsion between particles.
5. The average kinetic energy of the particles is proportional to temperature   
   – (in Kelvin!!!) – T ↑, KE ↑

* There is a distribution of speeds, some go faster than others   
  so overall there is an average kinetic energy of the sample.