Name: **Date:**   **Period:**   **Seat #:**

Write the definition of each concentration in terms of **solute**, **solvent**, and/or **solution**:

Molarity (M)	Molality (m)	Mole fraction (χ)	Weight percent (%)

Each of these concentrations involves grams or moles of solute, solvent, or solution. Determine those values.

Assume you dissolve 2.56 g of malic acid, $C_4H_6O_5$, in half a liter of water (500.0 g).	
Molarity of acid in solution	
Molality of acid in solution	
mole fraction of acid in solution	
weight percentage of acid in solution	

Fill in the blanks in the table. Aqueous solutions are assumed. Show all work

Compound	Molarity	Weight Percent	Mole Fraction
NaI	0.15		
C_2H_5OH		5.0	
$C_{12}H_{22}O_{11}$	0.15		