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| **Polar Covalent** |
| Partial charges on ends of molecule |
| Electrons shared unequally |
| Medium electronegativity difference between bonding elements |
| Nonmetal + nonmetal |
| Two bonding elements are not the same |
| Positively charged pole on end of molecule that is less electronegative |
| Negatively charged pole on end of molecule that is “hogging” the electrons |
| Water Molecule |
| Bond between S and O |
| Higher ionic character |
| Difference in charges at two ends is measured by dipole moment  Bonds in carbon dioxide |
| **Nonpolar Covalent** |
| No partial charges on ends of molecules |
| Electrons shared equally |
| Very low or no electronegativity difference between bonding elements |
| Nonmetal + nonmetal |
| Two bonding elements are same element |
| The molecule, carbon dioxide |
| Bond between Cl and Cl |
| Lower ionic character |