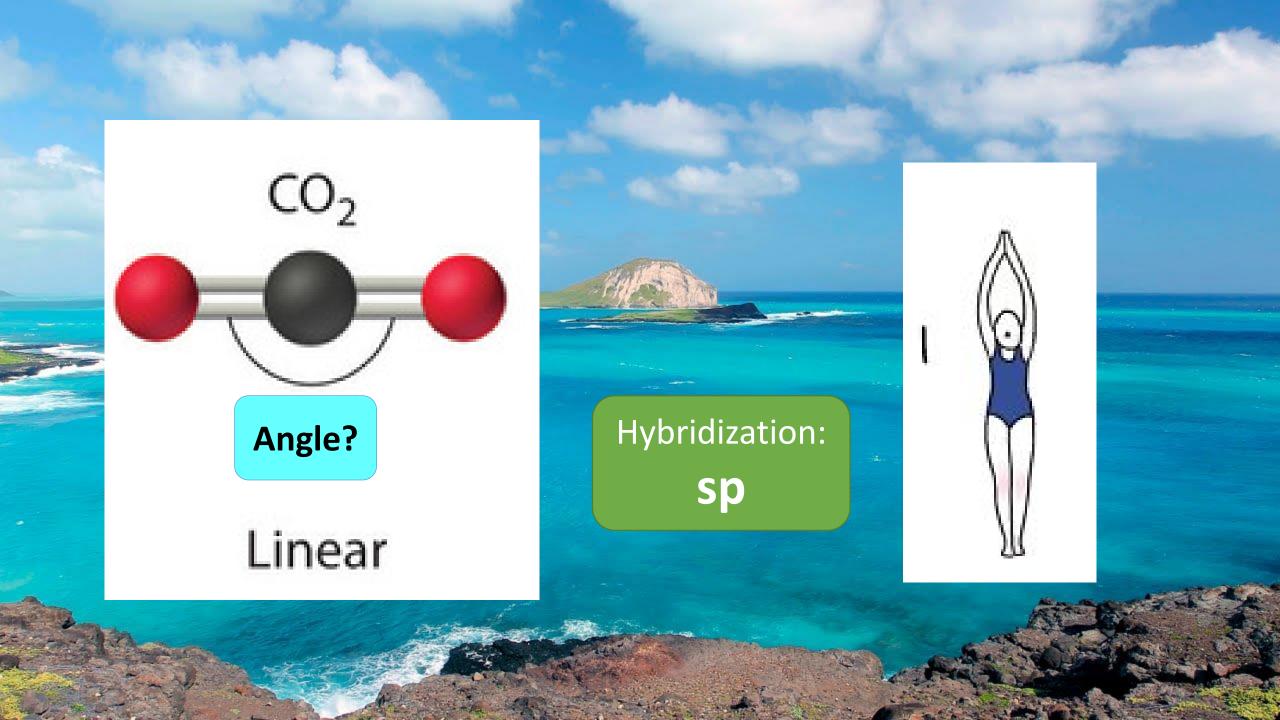
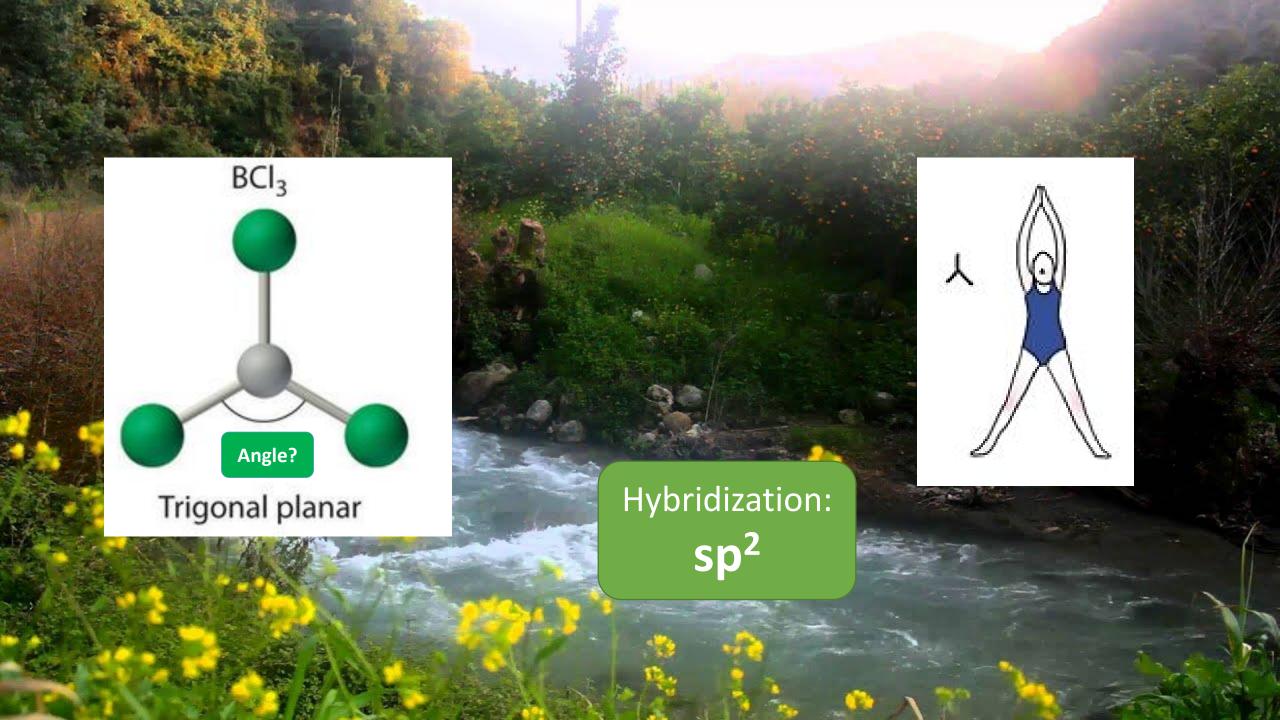
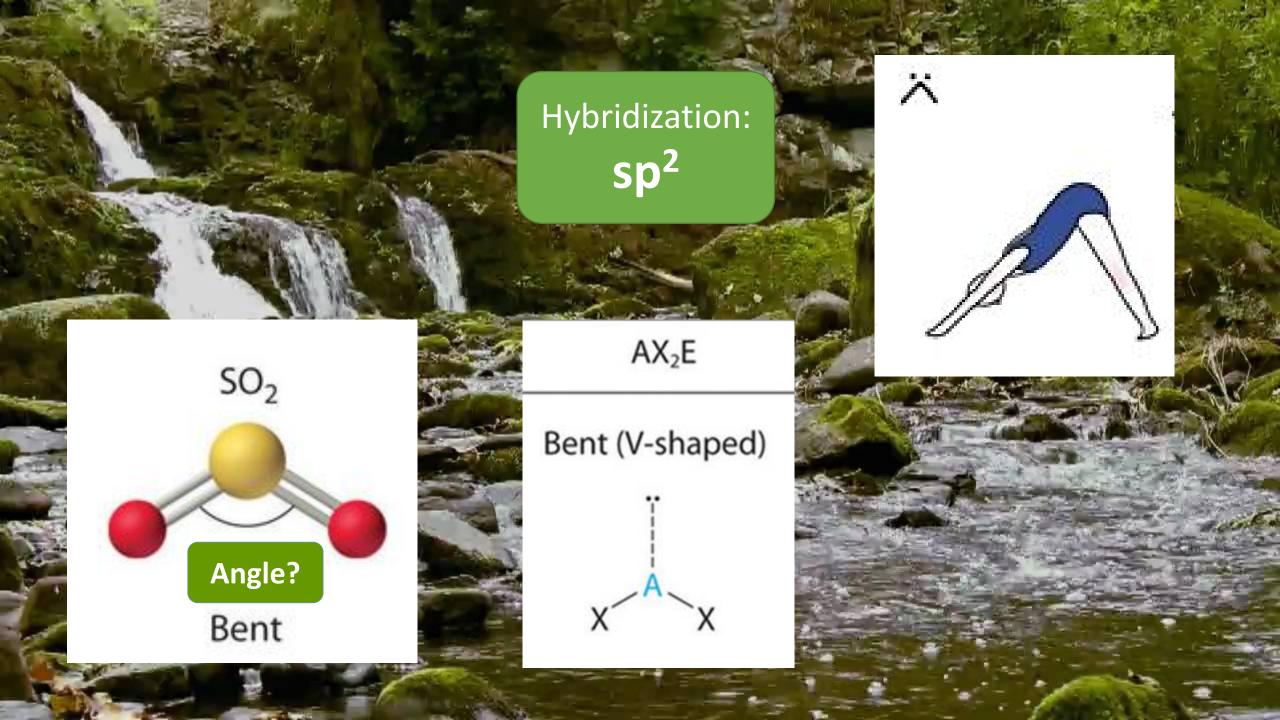
VSEPR YOGA

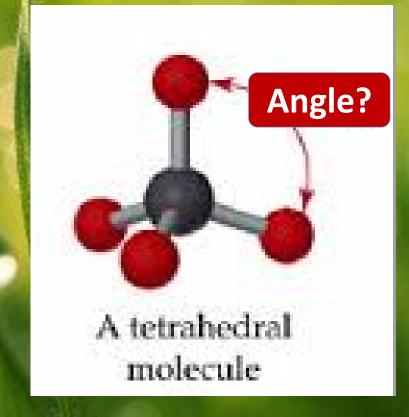
AP Chemistry Review Relaxation Day





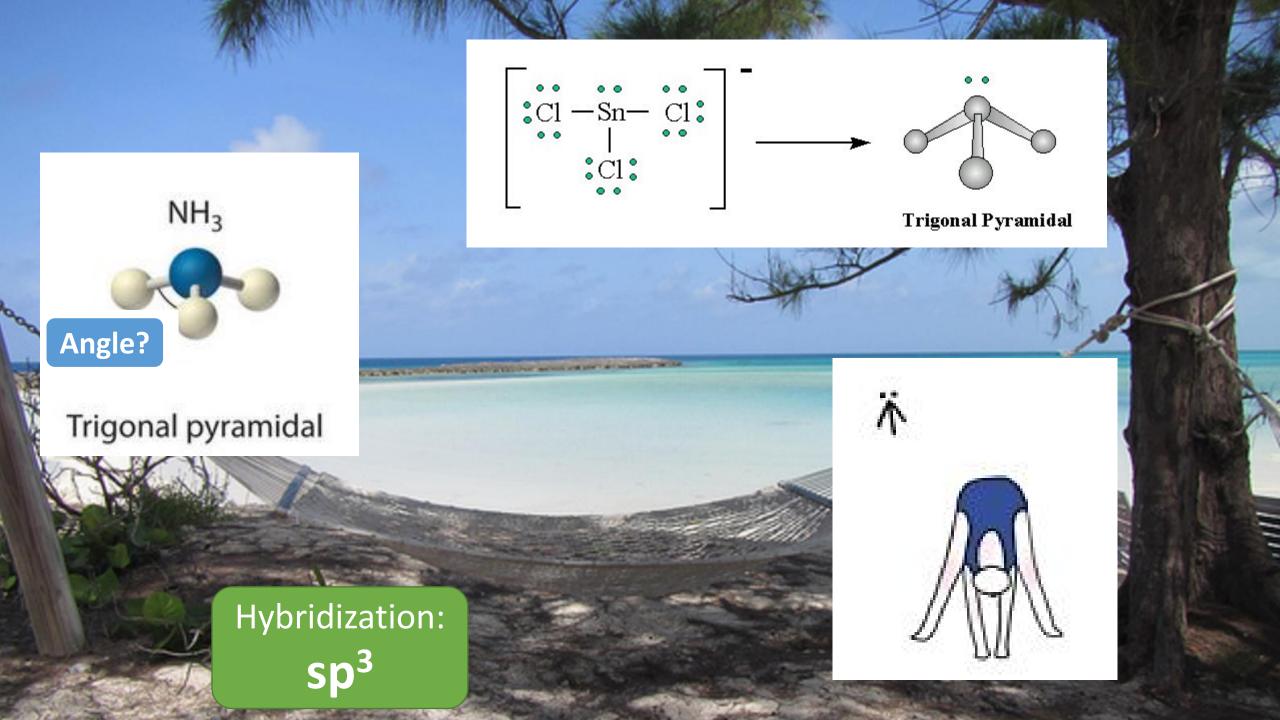


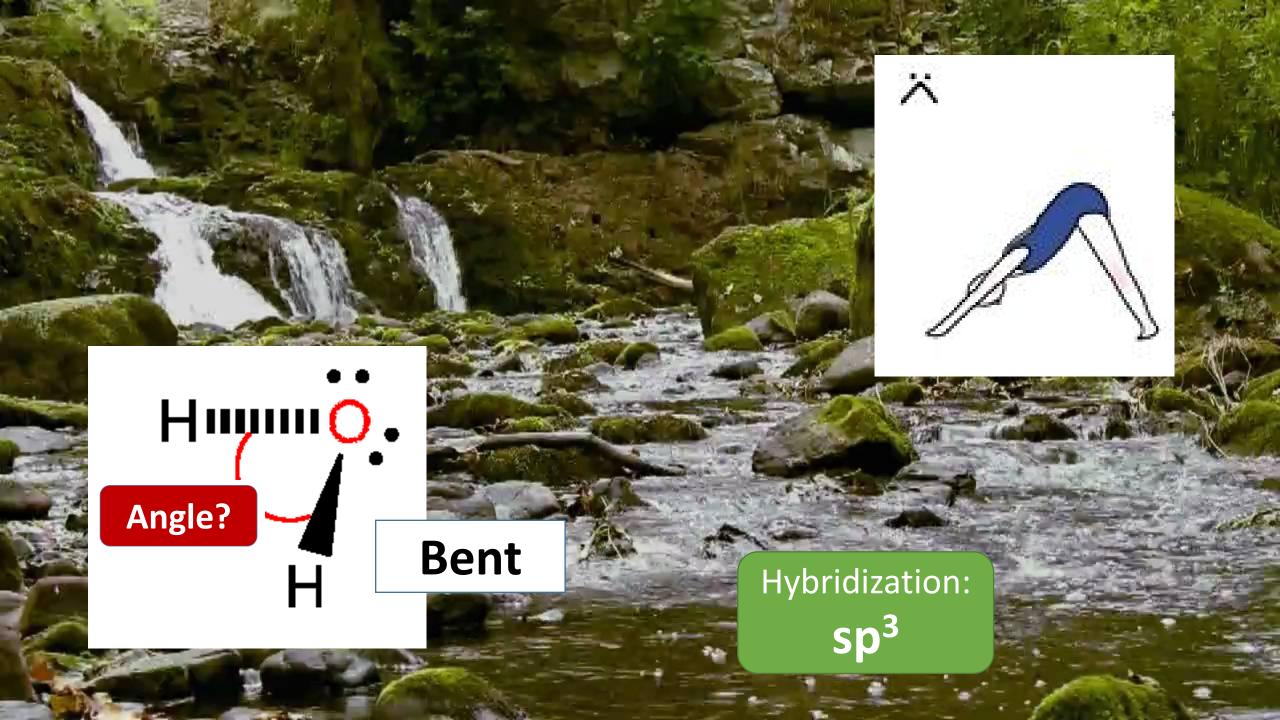
Hybridization: sp³

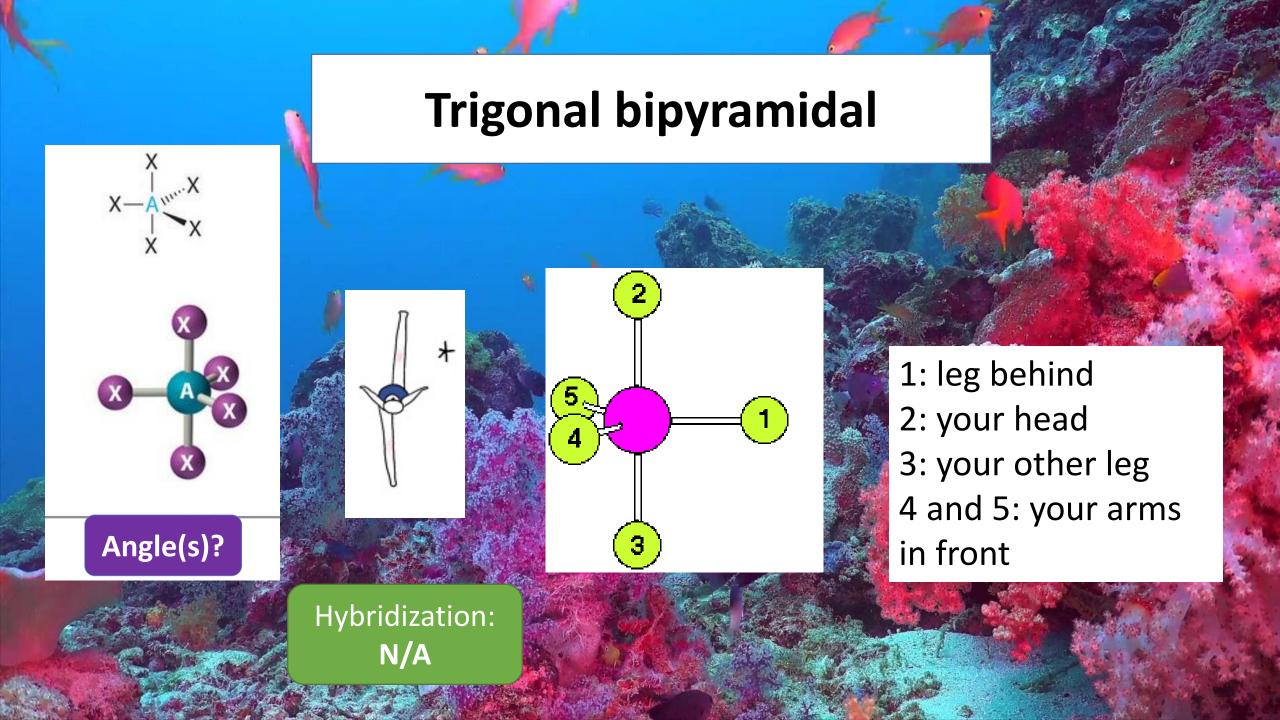




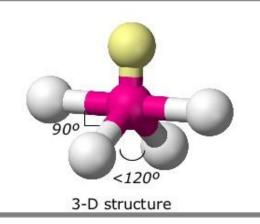


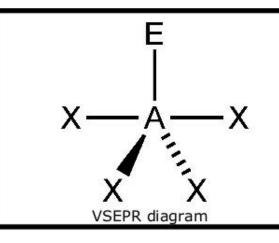






Seesaw



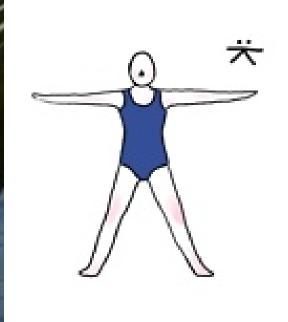


Example:

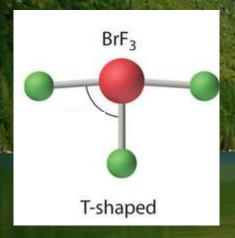
SF₄



Angle(s)?

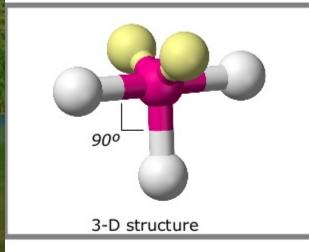


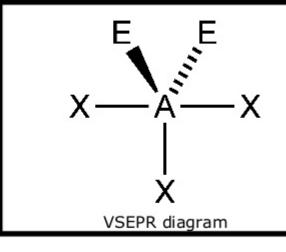
Hybridization: **N/A**



Angle?

T-shape

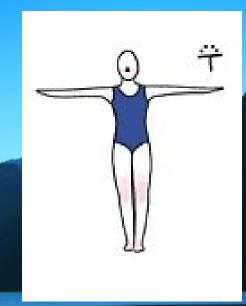




Examples: CIF₃, BrF₃



Hybridization: **N/A**







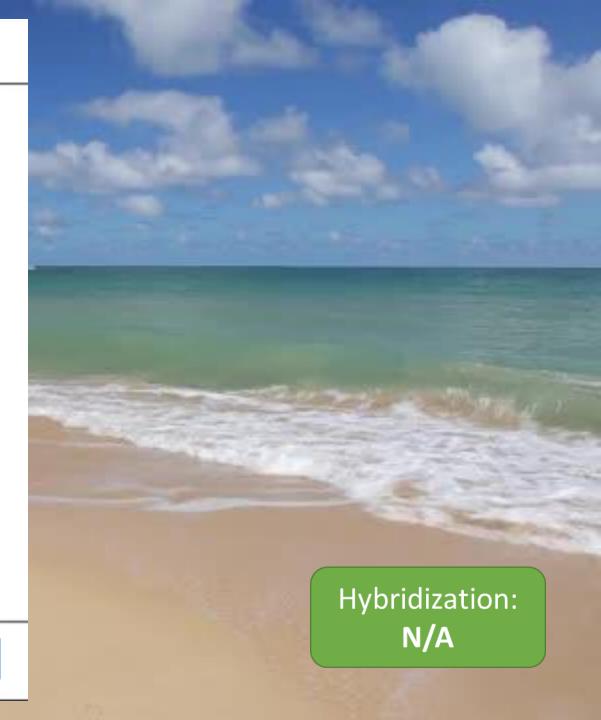


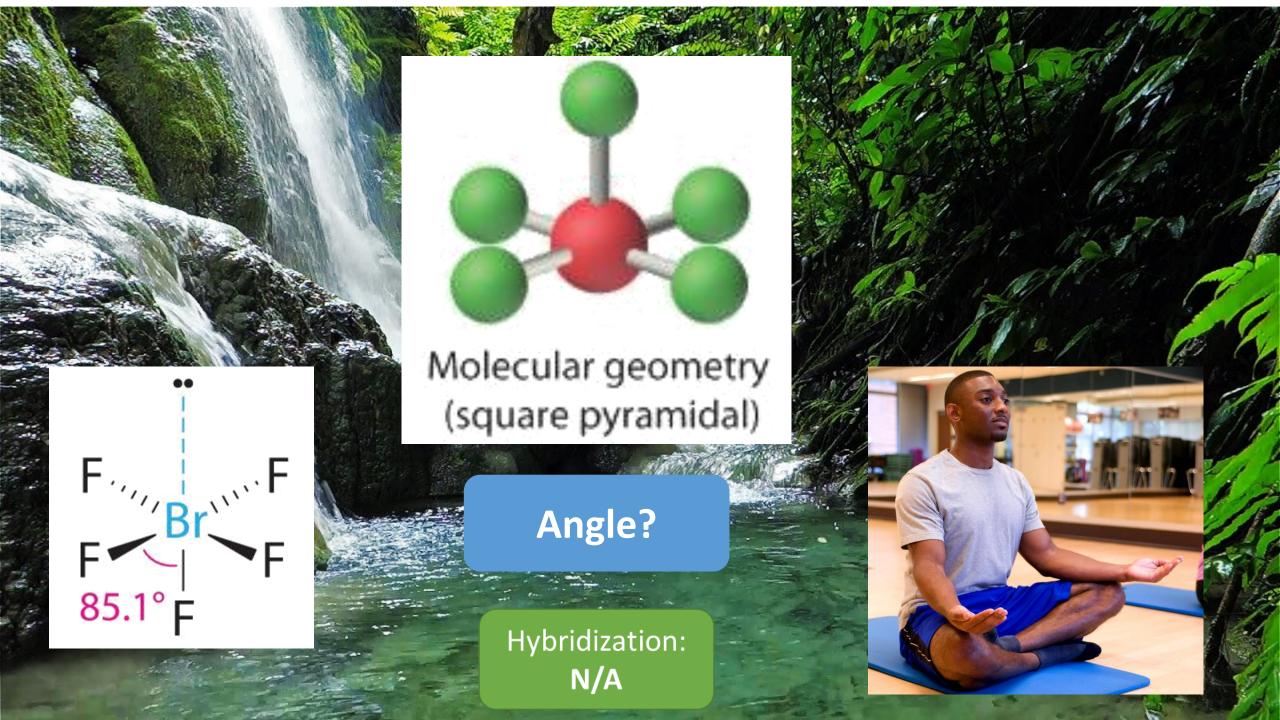
Octahedral

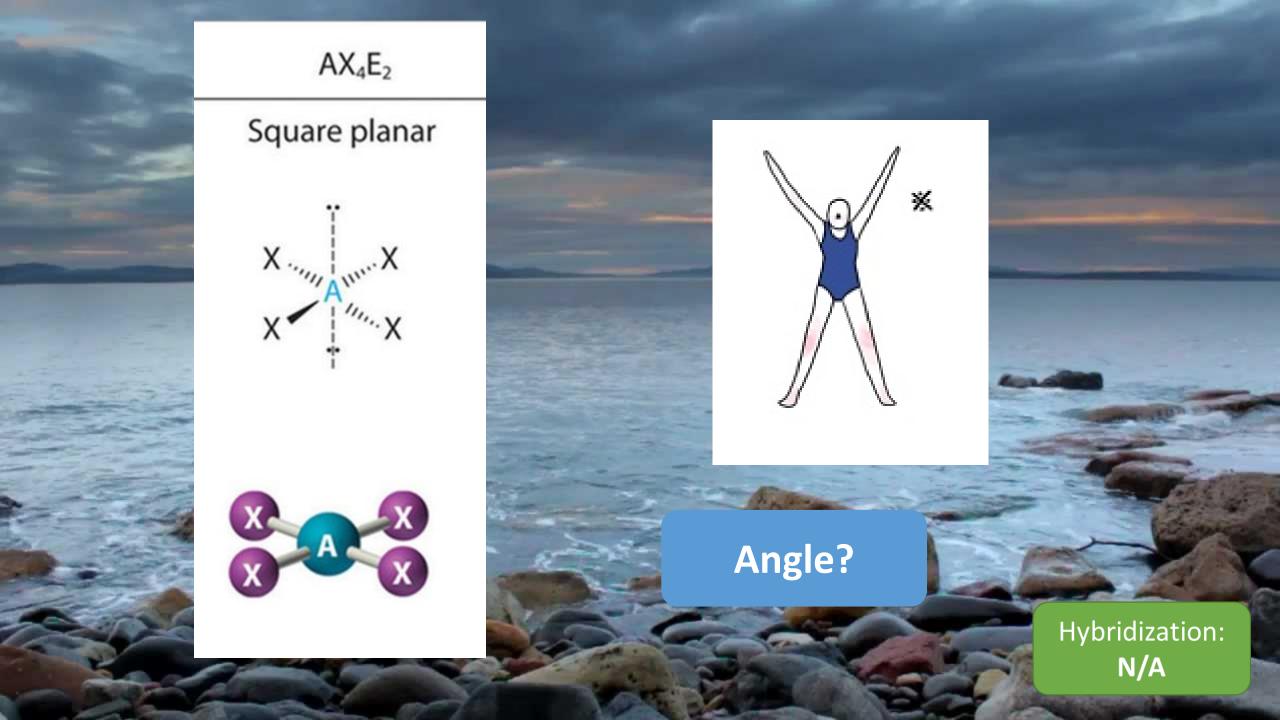




Angle?











Total no. of electron pairs = bond pairs + lone pairs	Basic Shape	Diagram	Bond angle
2	linear	0	18 O _o
3	trigonal Planar	trigonal planar	
4	tetrahedral		109.5"
5	trigonal bipyramidal	axial equatorial	120°
6	octahedral		90°

Number of Electron Dense Areas	Electron- Pair Geometry	Molecular Geometry				hybridization	
		No Lone Pairs	1 Ione Pair	2 lone Pairs	3 Ione Pairs	4 Ione Pairs	Hybridization
2	Linear	Linear					sp
3	Trigonal planar	Trigonal planar	Bent				sp ²
4	Tetrahedral	Tetrahedral	Trigonal pyramidal	Bent			sp ³
5	Trigonal bipyramidal	Trigonal bipyramidal	Sawhorse	T-shaped	Linear		N/A
6	Octahedral	Octahedral	Square pyramidal	Square planar	T-shaped	Linear	N/A