**More Intermolecular Forces**

*For questions 1-5, identify the main type of intermolecular force in each compound:*

*London dispersion force, dipole-dipole, hydrogen bond*

1) carbon disulfide

2) ammonia

3) oxygen

4) CH2F2

5) C2H6

*Rank the following compounds by increasing melting point – remember, as IMFs increase, the melting point increases*

6) C2H6, C2H5OH, C2H5I

7) H2S, H2O, H2

8) BBr3, BI3, BCl3

**More Intermolecular Forces - Key**

*For questions 1-5, identify the main type of intermolecular force in each compound:*

1) carbon disulfide

**London dispersion force**

2) ammonia

**Hydrogen bonding**

3) oxygen

**London dispersion force**

4) CH2F2

**Dipole-dipole forces**

5) C2H6

**London dispersion force**

*Rank the following compounds by increasing melting point:*

6) **C2H6, C2H5I, C2H5OH**

**LDF, DP-DP, HB**

7) **H2, H2S, H2O**

**LDF, DP-DP, HB**

8) **BCl3, BBr3, BI3**

**All nonpolar, only LDF, London forces increase as the size increases   
 (we count the number of electrons as a simplistic way of finding the size)**