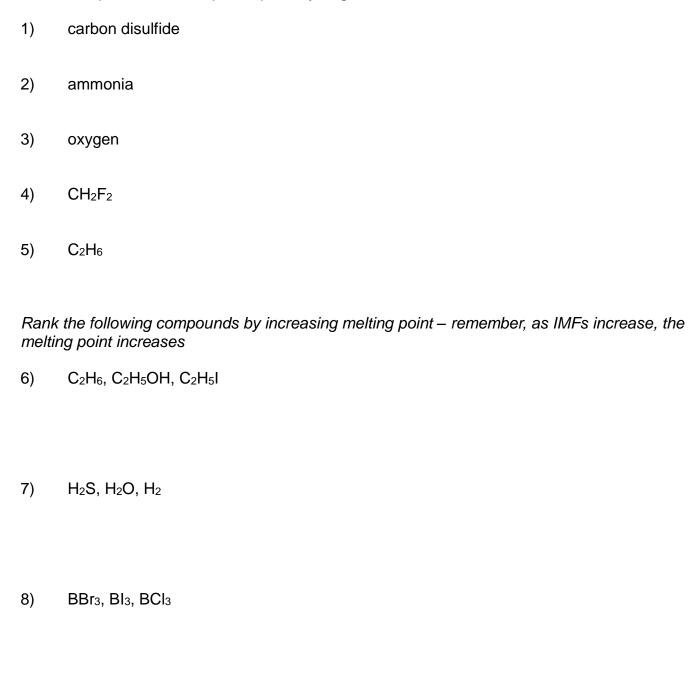
## **More Intermolecular Forces**

For questions 1-5, identify the main type of intermolecular force in each compound: London dispersion force, dipole-dipole, hydrogen bond



## More Intermolecular Forces - Key

For questions 1-5, identify the main type of intermolecular force in each compound:

- carbon disulfide
   London dispersion force
- 2) ammonia **Hydrogen bonding**
- 3) oxygen London dispersion force
- 4) CH<sub>2</sub>F<sub>2</sub> **Dipole-dipole forces**
- 5) C<sub>2</sub>H<sub>6</sub> London dispersion force

Rank the following compounds by increasing melting point:

- 6) C<sub>2</sub>H<sub>6</sub>, C<sub>2</sub>H<sub>5</sub>I, C<sub>2</sub>H<sub>5</sub>OH LDF, DP-DP, HB
- 7) H<sub>2</sub>, H<sub>2</sub>S, H<sub>2</sub>O LDF, DP-DP, HB
- 8) BCI<sub>3</sub>, BBr<sub>3</sub>, BI<sub>3</sub>
  All nonpolar, only LDF, London forces increase as the size increases
  (we count the number of electrons as a simplistic way of finding the size)