

PRACTICE PROBLEMS ON NET IONIC EQUATIONS

Show the complete ionic and net ionic forms of the following equations. If all species are spectator ions, please indicate that no reaction takes place. Note: you need to make sure the original equation is balanced before proceeding! Be sure to use your Solubility Rules chart for this exercise.

1. $\text{AgNO}_3(\text{aq}) + \text{KCl}(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{KNO}_3(\text{aq})$
2. $\text{Mg}(\text{NO}_3)_2(\text{aq}) + \text{Na}_2\text{CO}_3(\text{aq}) \rightarrow \text{MgCO}_3(\text{s}) + \text{NaNO}_3(\text{aq})$
3. strontium bromide(aq) + potassium sulfate(aq) \rightarrow strontium sulfate(s) + potassium bromide(aq)
4. manganese(II)chloride(aq) + ammonium carbonate(aq) \rightarrow manganese(II)carbonate(s) + ammonium chloride(aq)
5. chromium(III)nitrate(aq) + iron(II)sulfate(aq) \rightarrow chromium(III)sulfate(aq) + iron(II)nitrate(aq)

Please complete the following reactions, and show the complete ionic and net ionic forms of the equation:

6. $\text{K}_3\text{PO}_4(\text{aq}) + \text{Al}(\text{NO}_3)_3(\text{aq}) \rightarrow$
7. $\text{BeI}_2(\text{aq}) + \text{Cu}_2\text{SO}_4(\text{aq}) \rightarrow$
8. $\text{Ni}(\text{NO}_3)_3(\text{aq}) + \text{KBr}(\text{aq}) \rightarrow$
9. cobalt(III)bromide + potassium sulfide \rightarrow
10. barium nitrate + ammonium phosphate \rightarrow
11. calcium hydroxide + iron(III)chloride \rightarrow
12. rubidium fluoride + copper(II)sulfate \rightarrow