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| **Cathode rays have  identical properties regardless of element used** |  |
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| **Most of the particles passed right through**  **A few particles were greatly deflected.**  **Very few were GREATLY deflected** |  |

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| Dalton’s Atomic Theory (1808) | | |
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| 2 | **Atoms of a given element are identical in size, mass, and other properties** |  |
| 3 | **Atoms of different elements differ in size, mass, and other properties** |  |
| 4 | **Atoms cannot be subdivided, created, or destroyed** |  |
| 5 | **Atoms of different elements combine in simple whole-number ratios to form chemical compounds** |  |
| 6 | **In chemical reactions, atoms are combined, separated, or rearranged** |  |

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