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| **ICE Table Practice Problem #1**  If you have an initial concentration of [PCl5] at 1.3M, what are the concentrations of the products at equilibrium? Assume all reactants and products are aqueous and Keq = 78.3.  PCl5 🡪 PCl3 + Cl2   |  |  |  |  | | --- | --- | --- | --- | | **Rxn** | PCl5 🡪 PCl3 + Cl2 | | | | **I** |  |  |  | | **C** |  |  |  | | **E** |  |  |  | | **5%** |  |  |  | | **Answer** |  |  |  | |  | **ICE Table Practice Problem #1**  If you have an initial concentration of [PCl5] at 1.3M, what are the concentrations of the products at equilibrium? Assume all reactants and products are aqueous and Keq = 78.3.  PCl5 🡪 PCl3 + Cl2   |  |  |  |  | | --- | --- | --- | --- | | **Rxn** | PCl5 🡪 PCl3 + Cl2 | | | | **I** |  |  |  | | **C** |  |  |  | | **E** |  |  |  | | **5%** |  |  |  | | **Answer** |  |  |  | |
| **ICE Table Practice Problem #2**  In the following reaction, Keq = 9.3x10-7 at room temp. Calculate the equilibrium concentration of N2O4 in a flask initially containing only 3.00 M of NO2  2 NO2(g) 🡪 N2O4(g)   |  |  |  | | --- | --- | --- | | **Rxn** | 2 NO2(g) 🡪 N2O4(g) | | | **I** |  |  | | **C** |  |  | | **E** |  |  | | **5%** |  |  | | **Answer** |  |  | |  | **ICE Table Practice Problem #2**  In the following reaction, Keq = 9.3x10-7 at room temp. Calculate the equilibrium concentration of N2O4 in a flask initially containing only 3.00 M of NO2  2 NO2(g) 🡪 N2O4(g)   |  |  |  | | --- | --- | --- | | **Rxn** | 2 NO2(g) 🡪 N2O4(g) | | | **I** |  |  | | **C** |  |  | | **E** |  |  | | **5%** |  |  | | **Answer** |  |  | |

N-45

N-45