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| **ICE Table Practice Problem #1**If you have an initial concentration of [PCl5] at 1.3M, what are the concentrations of the products at equilibrium? Assume all reactants and products are aqueous and Keq = 78.3.PCl5 🡪 PCl3 + Cl2

|  |  |
| --- | --- |
| **Rxn** |  PCl5 🡪 PCl3 + Cl2 |
| **I** |  |  |  |
| **C** |  |  |  |
| **E** |  |  |  |
| **5%** |  |  |  |
| **Answer** |  |  |  |

 |  | **ICE Table Practice Problem #1**If you have an initial concentration of [PCl5] at 1.3M, what are the concentrations of the products at equilibrium? Assume all reactants and products are aqueous and Keq = 78.3.PCl5 🡪 PCl3 + Cl2

|  |  |
| --- | --- |
| **Rxn** |  PCl5 🡪 PCl3 + Cl2 |
| **I** |  |  |  |
| **C** |  |  |  |
| **E** |  |  |  |
| **5%** |  |  |  |
| **Answer** |  |  |  |

 |
| **ICE Table Practice Problem #2**In the following reaction, Keq = 9.3x10-7 at room temp. Calculate the equilibrium concentration of N2O4 in a flask initially containing only 3.00 M of NO22 NO2(g) 🡪 N2O4(g)

|  |  |
| --- | --- |
| **Rxn** | 2 NO2(g) 🡪 N2O4(g) |
| **I** |  |  |
| **C** |  |  |
| **E** |  |  |
| **5%** |  |  |
| **Answer** |  |  |

 |  | **ICE Table Practice Problem #2**In the following reaction, Keq = 9.3x10-7 at room temp. Calculate the equilibrium concentration of N2O4 in a flask initially containing only 3.00 M of NO22 NO2(g) 🡪 N2O4(g)

|  |  |
| --- | --- |
| **Rxn** | 2 NO2(g) 🡪 N2O4(g) |
| **I** |  |  |
| **C** |  |  |
| **E** |  |  |
| **5%** |  |  |
| **Answer** |  |  |

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N-45

N-45