

### ICE Table Practice Problem #1

If you have an initial concentration of  $[\text{PCl}_5]$  at 1.3M, what are the concentrations of the products at equilibrium? Assume all reactants and products are aqueous and  $K_{\text{eq}} = 78.3$ .



Rxn	$\text{PCl}_5$	$\rightarrow$	$\text{PCl}_3$	+	$\text{Cl}_2$
I					
C					
E					
5%					
Answer					

### ICE Table Practice Problem #1

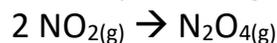
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Rxn	$\text{PCl}_5$	$\rightarrow$	$\text{PCl}_3$	+	$\text{Cl}_2$
I					
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E					
5%					
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### ICE Table Practice Problem #2

In the following reaction,  $K_{\text{eq}} = 9.3 \times 10^{-7}$  at room temp. Calculate the equilibrium concentration of  $\text{N}_2\text{O}_4$  in a flask initially containing only 3.00 M of  $\text{NO}_2$

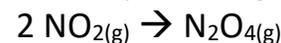


Rxn	$2 \text{NO}_{2(\text{g})}$	$\rightarrow$	$\text{N}_2\text{O}_{4(\text{g})}$
I			
C			
E			
5%			
Answer			

**N-45**

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I			
C			
E			
5%			
Answer			

**N-45**