

ICE Table Practice Problem #1

If you have an initial concentration of $[\text{PCl}_5]$ at 1.3M, what are the concentrations of the products at equilibrium? Assume all reactants and products are aqueous and $K_{\text{eq}} = 78.3$.



Rxn	PCl_5	\rightarrow	PCl_3	+	Cl_2
I					
C					
E					
5%					
Answer					

ICE Table Practice Problem #1

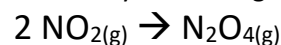
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Rxn	PCl_5	\rightarrow	PCl_3	+	Cl_2
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ICE Table Practice Problem #2

In the following reaction, $K_{\text{eq}} = 9.3 \times 10^{-7}$ at room temp. Calculate the equilibrium concentration of N_2O_4 in a flask initially containing only 3.00 M of NO_2

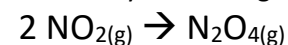


Rxn	$2 \text{NO}_{2(g)}$	\rightarrow	$\text{N}_2\text{O}_{4(g)}$
I			
C			
E			
5%			
Answer			

N-45

ICE Table Practice Problem #2

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5%			
Answer			

N-45