|  |  |  |
| --- | --- | --- |
| Naming AcidsBinary:Hydrogen + highly electronegative element1. Begins with **hydro**
2. Add the **root of the other element**
3. Add **–ic**
4. + **acid**

HBr - Hydrobromic acidHCl - Hydrochloric acidHI - Hydroiodic acidOxyacids:Hydrogen + oxygen + a third element1. Begins with **Root** **of ion** (not H or O) (sometimes starts with **per-** or **hypo-**)
2. Add **–ic, or -ous**
3. + **acid**

Names change a little depending on how many oxygens the anion comes with…Anion ends with **–ate** 🡪 change ending to **–ic**Anion ends with **–ite** 🡪 change ending to **–ous**Anion has **extra O than –ate** 🡪 start with **Per-**Anion has **fewer O than –ite** 🡪 start with **Hypo-** **ClO-** less O version 🡪 **Hypo**chlorous Acid**ClO2-** -ic version 🡪 Chlor**ous** Acid**ClO3-** -ate version 🡪 Chlor**ic** Acid**ClO4-** more O version 🡪 **Per**chloric Acid | Naming AcidsBinary:Hydrogen + highly electronegative element1. Begins with **hydro**
2. Add the **root of the other element**
3. Add **–ic**
4. + **acid**

HBr - Hydrobromic acidHCl - Hydrochloric acidHI - Hydroiodic acidOxyacids:Hydrogen + oxygen + a third element1. Begins with **Root** **of ion** (not H or O) (sometimes starts with **per-** or **hypo-**)
2. Add **–ic, or -ous**
3. + **acid**

Names change a little depending on how many oxygens the anion comes with…Anion ends with **–ate** 🡪 change ending to **–ic**Anion ends with **–ite** 🡪 change ending to **–ous**Anion has **extra O than –ate** 🡪 start with **Per-**Anion has **fewer O than –ite** 🡪 start with **Hypo-** **ClO-** less O version 🡪 **Hypo**chlorous Acid**ClO2-** -ic version 🡪 Chlor**ous** Acid**ClO3-** -ate version 🡪 Chlor**ic** Acid**ClO4-** more O version 🡪 **Per**chloric Acid | Naming AcidsBinary:Hydrogen + highly electronegative element1. Begins with **hydro**
2. Add the **root of the other element**
3. Add **–ic**
4. + **acid**

HBr - Hydrobromic acidHCl - Hydrochloric acidHI - Hydroiodic acidOxyacids:Hydrogen + oxygen + a third element1. Begins with **Root** **of ion** (not H or O) (sometimes starts with **per-** or **hypo-**)
2. Add **–ic, or -ous**
3. + **acid**

Names change a little depending on how many oxygens the anion comes with…Anion ends with **–ate** 🡪 change ending to **–ic**Anion ends with **–ite** 🡪 change ending to **–ous**Anion has **extra O than –ate** 🡪 start with **Per-**Anion has **fewer O than –ite** 🡪 start with **Hypo-** **ClO-** less O version 🡪 **Hypo**chlorous Acid**ClO2-** -ic version 🡪 Chlor**ous** Acid**ClO3-** -ate version 🡪 Chlor**ic** Acid**ClO4-** more O version 🡪 **Per**chloric Acid |
| 7 Strong Acids | 7 Strong Acids | 7 Strong Acids |
| 1. HCl – Hydrochloric Acid
2. HBr – Hydrobromic Acid
3. HI – Hydriodic Acid
 | 1. H2SO4 – Sulfuric Acid
2. HNO3 – Nitric Acid
3. HClO4 – Perchloric Acid
4. HClO3 – Chloric Acid
 | 1. HCl – Hydrochloric Acid
2. HBr – Hydrobromic Acid
3. HI – Hydriodic Acid
 | 1. H2SO4 – Sulfuric Acid
2. HNO3 – Nitric Acid
3. HClO4 – Perchloric Acid
4. HClO3 – Chloric Acid
 | 1. HCl – Hydrochloric Acid
2. HBr – Hydrobromic Acid
3. HI – Hydriodic Acid
 | 1. H2SO4 – Sulfuric Acid
2. HNO3 – Nitric Acid
3. HClO4 – Perchloric Acid
4. HClO3 – Chloric Acid
 |
| 8 Strong Bases | 8 Strong Bases | 8 Strong Bases |
| 1. LiOH – Lithium Hydroxide
2. NaOH – Sodium Hydroxide
3. KOH – Potassium Hydroxide
4. RbOH –Rubidium Hydroxide
5. CsOH – Cesium Hydroxide
 | 1. Ca(OH)2 – Calcium Hydroxide
2. Sr(OH)2 – Strontium Hydroxide
3. Ba(OH)2 – Barium Hydroxide
 | 1. LiOH – Lithium Hydroxide
2. NaOH – Sodium Hydroxide
3. KOH – Potassium Hydroxide
4. RbOH –Rubidium Hydroxide
5. CsOH – Cesium Hydroxide
 | 1. Ca(OH)2 – Calcium Hydroxide
2. Sr(OH)2 – Strontium Hydroxide
3. Ba(OH)2 – Barium Hydroxide
 | 1. LiOH – Lithium Hydroxide
2. NaOH – Sodium Hydroxide
3. KOH – Potassium Hydroxide
4. RbOH –Rubidium Hydroxide
5. CsOH – Cesium Hydroxide
 | 1. Ca(OH)2 – Calcium Hydroxide
2. Sr(OH)2 – Strontium Hydroxide
3. Ba(OH)2 – Barium Hydroxide

N-47 |

N-47

N-47