|  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- |
| **Weak Acid/Base Practice Problem #2**  You have 0.010 M NH3. Calculate the pH. Kb = 1.8 x 10-5  NH3 + H2O 🡨🡪 NH4+ + OH-   |  |  |  |  | | --- | --- | --- | --- | | **Rxn** | NH3 🡨🡪 NH4+ + OH- | | | | **I** |  |  |  | | **C** |  |  |  | | **E** |  |  |  | | **5%** |  |  |  | | **Answer** |  |  |  | |  | **Weak Acid/Base Practice Problem #2**  You have 0.010 M NH3. Calculate the pH. Kb = 1.8 x 10-5  NH3 + H2O 🡨🡪 NH4+ + OH-   |  |  |  |  | | --- | --- | --- | --- | | **Rxn** | NH3 🡨🡪 NH4+ + OH- | | | | **I** |  |  |  | | **C** |  |  |  | | **E** |  |  |  | | **5%** |  |  |  | | **Answer** |  |  |  | |
|  |  |  |
| **Weak Acid/Base Practice Problem #2**  You have 0.010 M NH3. Calculate the pH. Kb = 1.8 x 10-5  NH3 + H2O 🡨🡪 NH4+ + OH-   |  |  |  |  | | --- | --- | --- | --- | | **Rxn** | NH3 🡨🡪 NH4+ + OH- | | | | **I** |  |  |  | | **C** |  |  |  | | **E** |  |  |  | | **5%** |  |  |  | | **Answer** |  |  |  | |  | **Weak Acid/Base Practice Problem #2**  You have 0.010 M NH3. Calculate the pH. Kb = 1.8 x 10-5  NH3 + H2O 🡨🡪 NH4+ + OH-   |  |  |  |  | | --- | --- | --- | --- | | **Rxn** | NH3 🡨🡪 NH4+ + OH- | | | | **I** |  |  |  | | **C** |  |  |  | | **E** |  |  |  | | **5%** |  |  |  | | **Answer** |  |  |  | |