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| **Weak Acid/Base Practice Problem #1**  You have 1.00 M HOAc. Calc. the equilibrium concentrations of HOAc, H3O+, OAc-, and the pH if Ka = 1.8x10-5.  HOAc + H2O 🡨🡪 H3O + OAc-   |  |  |  |  | | --- | --- | --- | --- | | **Rxn** | HOAc 🡨🡪 H3O + OAc- | | | | **I** |  |  |  | | **C** |  |  |  | | **E** |  |  |  | | **5%** |  |  |  | | **Answer** |  |  |  | |  | **Weak Acid/Base Practice Problem #1**  You have 1.00 M HOAc. Calc. the equilibrium concentrations of HOAc, H3O+, OAc-, and the pH if Ka = 1.8x10-5.  HOAc + H2O 🡨🡪 H3O + OAc-   |  |  |  |  | | --- | --- | --- | --- | | **Rxn** | HOAc 🡨🡪 H3O + OAc- | | | | **I** |  |  |  | | **C** |  |  |  | | **E** |  |  |  | | **5%** |  |  |  | | **Answer** |  |  |  | |
|  |  |  |
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