## Weak Acid/Base Practice Problem #1
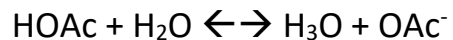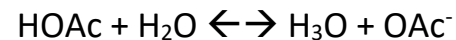
You have 1.00 M HOAc. Calc. the equilibrium concentrations of HOAc, $H_3O^+$, OAc$^-$, and the pH if Ka = $1.8 \times 10^{-5}$.

$$HOAc + H_2O \longleftrightarrow H_3O + OAc^-$$
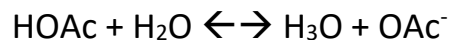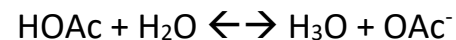
| Rxn | HOAc $\longleftrightarrow$ | $H_3O$ + | OAc$^-$ |
|---|---|---|---|
| I | | | |
| C | | | |
| E | | | |
| 5% | | | |
| Answer | | | |