	Makes the solution…	
Acidic + Neutral	Acidic	
Basic + Neutral	Basic	
Neutral + Neutral	Neutral	
Acidic + Basic	Compare Ka and Kb to determine which "wins"	
	$Ka_{(ion)} > Kb_{(ion)}$	Acidic
	$Ka_{(ion)} < Kb_{(ion)}$	Basic
	$Ka_{(ion)} = Kb_{(ion)}$	Neutral

Remember: Kw = Ka x Kb

$$Ka_{(acidic\ ion)} = \frac{Kw}{Kb\ (of\ where\ ion\ came\ from)}$$

$$Kb_{(basic\ ion)} = \frac{Kw}{Ka\ (of\ where\ ion\ came\ from)}$$

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