Steps to determine if a salt is acidic/basic/neutral

1. Identify ions that the salt came from
2. Determine if the ions will hydrolyze
   * Figure out if they came from a   
     strong or weak acid/base
     + From strong 🡪 ion won’t hydrolyze – neutral contribution
     + From weak 🡪 ion will hydrolyze –   
       acidic or basic contribution
3. If it hydrolyzes identify if the hydrolysis of the ion would form acid (H3O+)or base (OH-).
4. Figure out what the combo of each ion’s contribution would be to the solution
5. To determine the “winner” when acidic + basic
   * Compare the Ka and Kb values
   * The higher one means it is stronger, more dissociation so it will contribute more to the resulting solution

Steps to find the actual pH value of a salt solution

1. Do all the steps needed to determine which ion is the “strong” one – which one is being hydrolyzed?
2. Write the hydrolysis reaction for that ion (or ions)
3. ICE Table time! Yes! More ICE tables! They just won’t go away! ☺ Use your hydrolysis rxn for ICE Table
4. Find [H3O+] or [OH-] from ICE Tables
5. Continue on with normal pH type calculations using the concentrations you found from the ICE Table

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