Some molar masses to help speed things up for you ☺

C_3H_8 = 44.1

O_2 = 32

H_2O = 18

$Al_2(SO_3)_3$ = 294.2

NaOH = 40

Na_2SO_3 = 126.04

Al_2O_3 = 101.96

Fe = 55.85

$CuCl_2$ = 134.45

$NaNO_3$ = 84.995

$Cu(NO_3)_2$ = 187.56

NaCl = 58.443

$Ba_3(PO_4)_2$ = 601.92

$Na_3(PO_4)$ = 163.94

$BaCl_2$ = 208.2

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Some answers so you can check your work as you go ☺

1) O_2 is LR and C_3H_8 is XS
2) 0.0645 mol CO_2 made
3) 1.55 g H_2O made
4) 13.87 g C_3H_8 left

5) $Al_2(SO_3)_3$ is LR and NaOH is XS
6) 0.06798 mol $Al(OH)_3$ made
7) 12.85 g Na_2SO_3 made
8) 1.842 g NaOH left

9) Fe is LR and Al_2O_3 is XS
10) 0.163 mol Al made
11) 0.061 mol Fe_3O_4 made
12) 17.1 g Al_2O_3 left

13)
A) $CuCl_2 + 2NaNO_3 \rightarrow Cu(NO_3)_2 + 2NaCl$
B) $CuCl_2$ is LR
C) 0.224 mol NaCl made
D) 21.01 g $Cu(NO_3)_2$ made
E) 0.011 mol $NaNO_3$ left
F) 86.3% yield

14)
A) $6NaCl + Ba_3(PO_4)_2 \rightarrow 2Na_3(PO_4) + 3BaCl_2$
B) NaCl is LR
C) 935.0 g $Na_3(PO_4)$ made and 1781.2 g $BaCl_2$ made
D) 283.52 g $Ba_3(PO_4)_2$ left

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