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| **COMMENTS** | **SCORE** |
| **General*** Header info not filled out on every page
* Not neat/professional - other people cannot easily access the information
 | **5 points** |
| **Materials*** Missing some
* Not included
 | **5 points** |
| **Reagent Table*** Missing some info
* Not included
 | **5 points** |
| **Procedures*** Copied from lab sheet
* Missing key steps
 | **5 points** |
| **Observations/Data Table*** Missing either qualitative or quantitative data
* No title on data tables
* Titles not descriptive
* Too small/squished/messy
* Observations lacking detail or content
 | **15 points** |
| **Discussion Questions*** Questions copied from handout
* Questions not copied, but not paraphrased into your answer
* Missing Questions:
 | **35 points** |
| * Incorrect answers for Qs:
	+ 1
	+ 2
	+ 3
	+ 4
	+ 5
	+ 6
	+ 7
	+ 8
	+ 9
	+ 10
	+ 11
	+ 12
 | * Insufficient answers for Qs:
	+ 1
	+ 2
	+ 3
	+ 4
	+ 5
	+ 6
	+ 7
	+ 8
	+ 9
	+ 10
	+ 11
	+ 12
 | **1)** Electrons **2)** They went up to a higher energy level **3)** We gave them energy from the Bunsen burner **4)** They have different gaps in energy levels**5)** High energy = high frequency = short wavelengths **6)** High: purple, blue, green, yellow, orange, red :Low **7)** High: purple, blue, green, yellow, orange, red :Low**8)** High: red, orange, yellow, green, blue, purple :Low **9)** Copper – the color matched the Cu(NO3)2 and Cu(SO4) that we used**10)** We were not looking at them through a prism – the prism splits each individual wavelength up into lines. We just looked at the entire thing. **11)** If we used a prism to separate we might be able to. If we are just looking at the flame we wouldn’t be able to distinguish. **12)** Neon signs, lights, fireworks, the stars |
|  | **70 points TOTAL** |