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| Electron Configuration – an “address” for the electrons in an atom | | |
| **An Orbital is:** | **How do we describe orbitals?** | |
| **Different orbitals are in different energy levels** | **Different orbitals have different shapes** | |
| **Different orbitals have different orientations** | **Each orbital is only allowed to have two e-s** | |
| **Where do e- live? What is the address for one?**  State ------> Energy level  City ------> Type/shape of orbital Street ------> Orientation or orbital  House # ------> Spin up or spin down of e– | | **They can get REALLY long**  1s+½ , 1s-½ ,2s+½ , 2s-½  2px +½ , 2px -½ , 2py +½  2py -½ , 2pz +½ , 2pz -½  N-10 |

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| **Want to describe where ALL the e- in an  atom were?   Shrink it down and only list:**        **Example:** | | **Steps to finding all the electrons**   1. Pick an:  \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 2. Find the number of: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 3. Start putting electrons into the: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 4. Use an:  \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 5. List which:  \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ you used and   \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ electrons in each one | |
| **Rules for putting electrons in an orbital diagram:** | | | |
| 1. **Aufbau Principle**   *An electron occupies the lowest energy orbital that it can.*  Means: | 1. **Pauli Exclusion Principle**   *No two e-s in the same atom can have the same set of 4 quantum numbers*  Means: | | 1. **Hunds Rule**   *Orbitals of equal energy are each occupied by one e- before any orbital is occupied by a second e-.*  Means: |
| **Some Terms You Might Hear** | | | |

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