

Name: _____

Period: _____

Seat#: _____

On the blank chart below, you are to fill in the letters of the imaginary elements A – NN according to the clues below. Not all elements will have a clue. Use your knowledge of the periodic table and its organization to fill in the “elements.” Your “Periodic Table” represents periods 2-6 of the s and p blocks of the real periodic table. The “elements” below follow the same patterns that the actual elements on the periodic table follow.

	“s” block		“p” block					
Period 2								
Period 3								
Period 4								
Period 5								
Period 6								

Clues:

- The “elements” are grouped in the following Families:

A B C D E
F G H I J
K L M N O

P Q R S T
U V W X Y
Z AA BB CC DD

EE FF GG HH II
JJ KK LL MM NN

- Y** is the most electronegative element.
- GG** is the most reactive metal and the least electronegative element.
- If **K** lost 4 electrons, it would look like **E**, the smallest Noble Gas.
- If **I** gained 2 electrons, it would look like **A**, and if it had 2 less protons, it would be **K**.
- BB** is the biggest element in its family.
- AA** is smaller than **K** but bigger than **I**.
- CC** is the only element in its group that is diatomic.
- If **NN** lost 1 electron, its electrons would fill ONLY the s sublevel.
- JJ** is the second biggest element in its family.
- S** is in the same period as **JJ**.
- Z** and **L** are not in the same period.
- B** is the largest noble gas.
- If **II** were a real element, it would only have 3 electrons.
- FF** is in the same period with **U** but not **O**.
- H** is to **JJ** as **I** is to **KK**.
- U** has a larger ionization energy than **I**.
- M**’s valence electrons are in the 4th energy level.
- The shielding effect for **EE** is greater than for **HH**.
- F** is the smallest element in the group that would form charges of -2.
- G** gains 2 electrons to look like **B**.
- P**’s atomic number is 1 more than **GG**.
- T** is not in the same period with **Z**.
- L** is not in the same period as **GG**, but is in the same period with **C**.
- When **W** gains 1 electron, it’s a 4p⁶.
- R** has 1 less energy level than **M**.
- The valence electrons for **X** and **LL** are not in the same energy level.
- MM** and **Z** are in the same period.

