**Name: Period: Seat#:**

**Worksheet #7**

**Review your Periodic Table Structure knowledge:**

|  |  |  |
| --- | --- | --- |
| 1. Where are the most active metals located?
 | 1. Where are the most active non-metals located?
 | 1. As you go across a period →, does the atomic size decrease or increase? Why?
 |
| 1. As you travel down a group, does the atomic size decrease or increase? Why?
 | 1. Is a negative ion larger or smaller than its parent atom?
 | 1. Is a positive ion larger or smaller than its parent atom?
 |
| 1. As you go from left to right across a period, does the first ionization energy generally decrease or increase? Why?
 | 1. As you go down a group, does the first ionization energy generally decrease of increase? Why?
 | 1. Where is the highest electronegativity found?
 |
| 1. Where is the lowest electronegativity found?
 | 1. Elements of Group 1A are called:
 | 1. Elements of Group 2A are called:
 |
| 1. Elements in the middle of the periodic table are called:
 | 1. Group 7A elements are called:
 | 1. Group 8A elements are called:
 |
| 1. From left to right across the periodic table, do the elements go from metals to nonmetals, or nonmetals to metals?
 | 1. The most active element in Group 7A is:
 | 1. What type of orbitals are filling across the Transition Elements?
 |
| 1. Elements within a group have the same number of what?
 | 1. Are the majority of elements in the periodic table metals or nonmetals?
 | 1. Elements in the periodic table are arranged according to their what?
 |

**Rank the atoms from smallest to largest atomic radius**

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| 1. Li, C, F
 | 1. Li, Na, K
 | 1. Ge, P, O
 | 1. C, N, Al
 | 1. Al, Cl, Cu
 |

**Rank the atoms from lowest to highest ionization energy.**

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| 1. Mg, Si, S
 | 1. Mg, Ca, Ba
 | 1. F, Cl, Br
 | 1. Ba, Cu, Ne
 | 1. Si, P, He
 |

**Rank the atoms from lowest to highest electronegativity energy.**

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| 1. Li, C, N
 | 1. Ne, C, O
 | 1. Si, P, O
 | 1. Mg, K, P
 | 1. S, F, He
 |

**Rank the atoms from smallest to largest electron affinity**

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| 1. Li, C, F
 | 1. Li, Na, K
 | 1. Ge, P, O
 | 1. C, N, Al
 | 1. Al, Cl, Cu
 |

**Circle the correct element.**

Li Si S metal

N P As smallest ionization energy

K Ca Sc largest atomic mass

S Cl Ar member of the halogen family

Al Si P greatest electronegativity

Ga Al Si largest atomic radius

V Nb Ta largest atomic number

Te I Xe member of noble gases

Si Ge Sn 4 energy levels

Li Be B member of alkali metals

As Se Br 6 valence electrons

H Li Na nonmetal

Hg Tl Pb member of transition metals

Na Mg Al electron config. ending in s2p1

Pb Bi Po metalloid

B C N gas at room temperature

Ca Sc Ti electron config. ending in s2d2