**Worksheet #9**

 **Name: Period: Seat#:**

**Answer the following questions:**

|  |  |
| --- | --- |
| 1. What are the three types of bonds and how are their electron positions different?
 | 1. Why do you need to use prefixes for naming covalent bonds and not for naming ionic bonds?
 |
| 1. Why does carbon dioxide have two double bonds?
 | 1. Why can some elements have more than 8 electrons in their valance shell and what do we call it when they do?
 |
| 1. List the Roman numerals from 1 to 10.
 |

**Complete the following table:**

|  |  |  |
| --- | --- | --- |
| **Formula**  | **Type of Bond** | **Name** |
| 1. **Na2SO4**
 |  |  |
| 1. **SiO2**
 |  |  |
|  |  | **Lead (II) nitrite**  |
|  |  | **Chromium (III) oxide**  |
| 1. **HgO**
 |  |  |
|  |  | **Iron (II) phosphate**  |
|  |  | **Hexaboron silicide**  |
| 1. **SCl4**
 |  |  |
| 1. **P4S5**
 |  |  |
| 1. **NaHCO3**
 |  |  |

**Draw the Lewis Structure for the following molecules:**

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| **Molecule** | **Lewis Structure** | **Description** | **Molecule** | **Lewis Structure** | **Description** |
| **SF6** |  | # of Single Bonds | # of Double Bonds | **Sulfate ion**  |  | # of Single Bonds | # of Double Bonds |
| # Valence electrons | # of Triple Bonds | # ofLone Pairs | # Valence electrons | # of Triple Bonds | # ofLone Pairs |
|  **CH3OH** |  | # of Single Bonds | # of Double Bonds |  **BFCl2** |  | # of Single Bonds | # of Double Bonds |
| # Valence electrons | # of Triple Bonds | # ofLone Pairs | # Valence electrons | # of Triple Bonds | # ofLone Pairs |
| 1. **O3**
 |  | # of Single Bonds | # of Double Bonds | 1. **BeH2**
 |  | # of Single Bonds | # of Double Bonds |
| # Valence electrons | # of Triple Bonds | # ofLone Pairs | # Valence electrons | # of Triple Bonds | # ofLone Pairs |
| 1. **SiI4**
 |  | # of Single Bonds | # of Double Bonds | 1. **K2SO3**
 |  | # of Single Bonds | # of Double Bonds |
| # Valence electrons | # of Triple Bonds | # ofLone Pairs | # Valence electrons | # of Triple Bonds | # ofLone Pairs |
|  **Fe3(PO4)2** |  | # of Single Bonds | # of Double Bonds |  **NaOH** |  | # of Single Bonds | # of Double Bonds |
| # Valence electrons | # of Triple Bonds | # ofLone Pairs | # Valence electrons | # of Triple Bonds | # ofLone Pairs |