**Name: Period: Seat#:**

**Worksheet #3B**

**Directions:** These are the Qs that went with WS #3. For any questions you did not complete in class, use this paper to finish the problems on WS #3. Put your answers on WS #3 not this paper, this is just the questions!

1. What charge will Nitrogen have when it forms an ion? What will its name be?
2. What is the formula for Calcium Hydroxide?
3. Mg(OH)2 Name and type of bond?
4. Ba(NO3)2 Name and type of bond?
5. What is the formula for lithium sulfate?
6. What charge will sulfur have when it forms and ion? What will its name be?
7. CoF3 Name and type of bond?
8. What is the formula for Aluminum sulfate?
9. What is the formula for Iron (III) Chloride?
10. Fe2O3 Name and type of bond?
11. Co3N2 Name and type of bond?
12. \_\_\_Mg + \_\_\_ N2 🡪\_\_\_Mg3N2
13. \_\_\_KOH + H3PO4 🡪 \_\_\_K3PO4 + \_\_\_H2O
14. \_\_\_Fe + \_\_\_H2SO4 🡪 \_\_\_Fe2 (SO4)3 + \_\_\_H2
15. \_\_\_C2H6 +\_\_\_O2🡪 \_\_\_H2O +\_\_\_CO2
16. Identify type of reaction
3 Mg + N2 🡪Mg3N2
17. Identify type of reaction
3 KOH + H3PO4 🡪 K3PO4 + 3 H2O
18. Identify type of reaction
2 Fe +3 H2SO4 🡪 Fe2 (SO4)3 + 3 H2
19. Write and Balance:
Sodium plus Oxygen yield Sodium Oxide
20. Write and Balance: Potassium oxide is added to water producing Potassium Hydroxide
21. Write and Balance: Solid Carbon (graphite) is added oxygen makes carbon dioxide
22. Write and Balance: Beryllium oxide is added to carbon dioxide producing Beryllium Carbonate
23. Write and Balance: Aluminum oxide added to water makes aluminum hydroxide
24. Write and Balance: \_\_Zn+ \_\_HCl →
25. Write and Balance:
 \_\_Al2(SO4)3(aq) + \_\_Ca3(PO4)2(aq) →
26. Write and Balance: \_\_\_C7H6O + \_\_\_O2 →
27. Write and Balance: Barium metal reacts with Iron (III) sulfate to produce…
28. Write and Balance: The combustion of decane. (C10H22)…
29. Write and Balance: A solution of Potassium Carbonate plus a solution of barium chloride produces…
30. Determine if a precipitate forms when the following solutions are mixed. Write the overall equation, the total ionic equation, and the net ionic equation. A solution of sodium fluoride is mixed with a solution of barium nitrate.
31. Determine if a precipitate forms when the following solutions are mixed. Write the overall equation, the total ionic equation, and the net ionic equation. A solution of potassium bromide is mixed with a solution of ammonium sulfate.