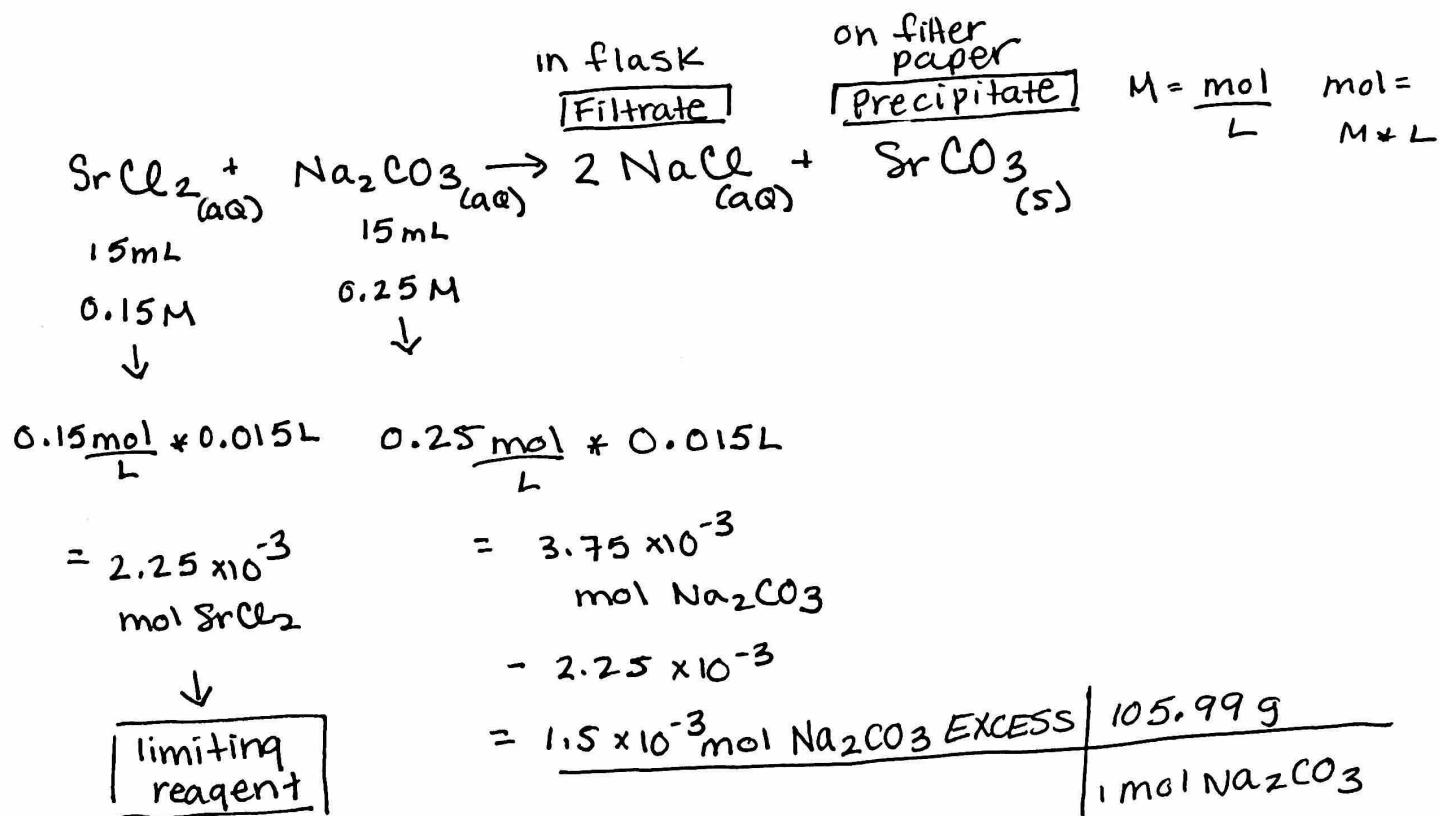




Theoretical yield

= 0.159 g extra Na₂CO₃ in filter flask

2.25 x 10 ⁻³ mol SrCl ₂	2 mol NaCl	58.4 g NaCl	= 0.292 g NaCl
1 mol SrCl ₂	1 mol NaCl		

2.25 x 10 ⁻³ mol SrCl ₂	1 mol SrCO ₃	147.6 g SrCO ₃	= 0.369 g SrCO ₃
1 mol SrCl ₂	1 mol SrCO ₃		

Actual yield

NaCl = 0.457g → 157% yield b/c SrCO₃ & excess reagent contaminated it

$$0.04797\text{g SrCO}_3 + 0.159\text{g Na}_2\text{CO}_3 + 0.292\text{g NaCl} = 0.499\text{g}$$

actual of everything

SrCO₃ = 0.321g → 87% yield b/c lost some to under filter paper