|  |  |  |  |
| --- | --- | --- | --- |
| Total pressure of mixture  (3.0 mol He and 4.0 mol Ne) is 97.4 kPa. *Find the partial pressure of each gas* | Total pressure of mixture  (3.0 mol He and 4.0 mol Ne) is 97.4 kPa. *Find the partial pressure of each gas* | Total pressure of mixture  (3.0 mol He and 4.0 mol Ne) is 97.4 kPa. *Find the partial pressure of each gas* | Total pressure of mixture  (3.0 mol He and 4.0 mol Ne) is 97.4 kPa. *Find the partial pressure of each gas* |
| 80.0 g each of He, Ne, and Ar are in a container. The total pressure is 780 mm Hg. Find each gas’s partial pressure. | 80.0 g each of He, Ne, and Ar are in a container. The total pressure is 780 mm Hg. Find each gas’s partial pressure. | 80.0 g each of He, Ne, and Ar are in a container. The total pressure is 780 mm Hg. Find each gas’s partial pressure. | 80.0 g each of He, Ne, and Ar are in a container. The total pressure is 780 mm Hg. Find each gas’s partial pressure. |
| Two 1.0 L containers, A and B, contain gases under 2.0 and 4.0 atm, respectively. Both gases are forced into Container C (w/vol. 2.0 L). Find total pres. of mixture in C. | Two 1.0 L containers, A and B, contain gases under 2.0 and 4.0 atm, respectively. Both gases are forced into Container C (w/vol. 2.0 L). Find total pres. of mixture in C. | Two 1.0 L containers, A and B, contain gases under 2.0 and 4.0 atm, respectively. Both gases are forced into Container C (w/vol. 2.0 L). Find total pres. of mixture in C. | Two 1.0 L containers, A and B, contain gases under 2.0 and 4.0 atm, respectively. Both gases are forced into Container C (w/vol. 2.0 L). Find total pres. of mixture in C. |
| Find total pressure of mixture in Container D. | Find total pressure of mixture in Container D. | Find total pressure of mixture in Container D. | Find total pressure of mixture in Container D. |
| https://sites.google.com/site/mohshchemwithmrsp/_/rsrc/1472859898453/unit-4---gas-laws/dalton-s-law/daltons%20law.png?height=155&width=320 | https://sites.google.com/site/mohshchemwithmrsp/_/rsrc/1472859898453/unit-4---gas-laws/dalton-s-law/daltons%20law.png?height=155&width=320 | https://sites.google.com/site/mohshchemwithmrsp/_/rsrc/1472859898453/unit-4---gas-laws/dalton-s-law/daltons%20law.png?height=155&width=320 | https://sites.google.com/site/mohshchemwithmrsp/_/rsrc/1472859898453/unit-4---gas-laws/dalton-s-law/daltons%20law.png?height=155&width=320 |
| Hydrogen gas is collected over water at 22°C. Find the pressure of the dry gas if the atmospheric pressure is 708 mmHg. | Hydrogen gas is collected over water at 22°C. Find the pressure of the dry gas if the atmospheric pressure is 708 mmHg. | Hydrogen gas is collected over water at 22°C. Find the pressure of the dry gas if the atmospheric pressure is 708 mmHg. | Hydrogen gas is collected over water at 22°C. Find the pressure of the dry gas if the atmospheric pressure is 708 mmHg. |
| A gas is collected over water at a temp of 35°C while the barometric pressure is 0.976 atm. What is the partial pressure of the dry gas? | A gas is collected over water at a temp of 35°C while the barometric pressure is 0.976 atm. What is the partial pressure of the dry gas? | A gas is collected over water at a temp of 35°C while the barometric pressure is 0.976 atm. What is the partial pressure of the dry gas? | A gas is collected over water at a temp of 35°C while the barometric pressure is 0.976 atm. What is the partial pressure of the dry gas? |

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