

Solubility of Some Ionic Compounds in Water		
<i>Always Soluble</i> Alkali metals (Li^+ , Na^+ , K^+ , Rb^+ , Cs^+), NH_4^+ Nitrate = NO_3^- , Chlorate = ClO_3^- , Perchlorate = ClO_4^- , Acetate = $\text{C}_2\text{H}_3\text{O}_2^-$		mnemonics
<i>Generally soluble</i> Cl^- , Br^- , I^- Soluble except: Ag^+ , Pb^{2+} , Hg_2^{2+} F^- Soluble except: Ca^{2+} , Sr^{2+} , Ba^{2+} , Pb^{2+} , Mg^{2+}		AP/H CBS-PM
SO_4^{2-} Soluble except: Ca^{2+} , Sr^{2+} , Ba^{2+} , Pb^{2+}		CBS-PBS
<i>Generally Insoluble</i> O^{2-} , OH^- Insoluble except: Alkali metals and NH_4^+ Ca^{2+} , Sr^{2+} , Ba^{2+} somewhat soluble CO_3^{2-} , PO_4^{3-} , S^{2-} , SO_3^{2-} , $\text{C}_2\text{O}_4^{2-}$, CrO_4^{2-} Insoluble except: Alkali metals and NH_4^+		CBS

Not Soluble = forms precipitate

Soluble = dissolves in water (aqueous)