|  |  |
| --- | --- |
| **Solubility of Some Ionic Compounds in Water** | |
| *Always Soluble*  Alkali metals (Li+, Na+, K+, Rb+, Cs+), NH4+  Nitrate = NO , Chlorate = ClO , Perchlorate = ClO , Acetate =  3 3 4  C2H3O   2 | mnemonics |
| *Generally soluble*  Cl, Br-, I Soluble except: Ag+, Pb2+, Hg2  2+  F Soluble except: Ca2+, Sr2+, Ba2+, Pb2+, Mg2+ | AP/H CBS-PM |
| 2 2+ 2+ 2+ 2+  SO4 Soluble except: Ca , Sr , Ba , Pb | CBS-PBS |
| *Generally Insoluble*  O2, OH Insoluble except: Alkali metals and NH +  4  Ca2+, Sr2+, Ba2+ somewhat soluble  CO 2, PO 3, S2, SO 2, C O 2, CrO 2 3 4 3 2 4 4  Insoluble except: Alkali metals and NH +  4 | CBS |

**Not Soluble** = forms precipitate **Soluble** = dissolves in water (aqueous)