|  |
| --- |
| **Solubility of Some Ionic Compounds in Water** |
| *Always Soluble*Alkali metals (Li+, Na+, K+, Rb+, Cs+), NH4+Nitrate = NO , Chlorate = ClO , Perchlorate = ClO , Acetate =3 3 4C2H3O 2 | mnemonics |
| *Generally soluble*Cl, Br-, I Soluble except: Ag+, Pb2+, Hg22+F Soluble except: Ca2+, Sr2+, Ba2+, Pb2+, Mg2+ | AP/H CBS-PM |
| 2 2+ 2+ 2+ 2+SO4 Soluble except: Ca , Sr , Ba , Pb | CBS-PBS |
| *Generally Insoluble*O2, OH Insoluble except: Alkali metals and NH +4Ca2+, Sr2+, Ba2+ somewhat solubleCO 2, PO 3, S2, SO 2, C O 2, CrO 2 3 4 3 2 4 4Insoluble except: Alkali metals and NH +4 | CBS |

**Not Soluble** = forms precipitate **Soluble** = dissolves in water (aqueous)