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| SET #1 **Questions #1 - 8** |  | SET #2 **Questions #9 - 16** |
| *Answer #1*  **Strontium Oxide** |  | *Answer #9*  **Single replacement/single displacement** |
| *Answer #2*  **421.61 g/mol** |  | *Answer #10*  **CI2** |
| *Answer #3*  **Sea of electrons, delocalized electrons etc.** |  | *Answer #11* |
| *Answer #4*  **Covalent bond** |  | *Answer #12*  **tetrahedral** |
| *Answer #5*  **2 mol O2 : 1 mol CH4** |  | *Answer #13*  **68.15 g/mol** |
| *Answer #6*  **Hydrogen bonding**  **Remember H-NOF!** |  | *Answer #14*  **trigonal pyramidal** |
| *Answer #7*  **Metallic < Ionic Lattice  < Network Covalent** |  | *Answer #15* |
| *Answer#8*  **Cu(SO4)2** |  | *Answer #16*  **Combustion!** |
| SET #3 **Questions #17 - 25** |  | SET #4 **Questions #26 - 33** |
| *Answer #17*  **AlPO4 + 3Rb(NO2) →  Al(NO2)3 + Rb3(PO4)** |  | *Answer #26*  **Na2 CO3** |
| *Answer #18*  **Unequally!** |  | *Answer #27*  **106g/mol** |
| *Answer #19*  **10 mol ZnO!**  **Same molar ratio!** |  | *Answer #28*  **Fe2(SO4)3= 400.1 g/mole** |
| *Answer #20*  **31.1g ZnO** |  | *Answer #29*  **0.111 moles  Fe2(SO4)3** |
| *Answer #21*  **H2 F2 Cl2 Br2  I2 O2 N2** |  | *Answer #30*  **1087.7g  KCl** |
| *Answer #22*  **77.98 g/mol** |  | *Answer #31*  **1592.68 g Fe2(CO3)3** |
| *Answer #23*  **L of A→molA→mol B→L of B**  **76L H2O** |  | *Answer #32*  **Double displacement.** |
| *Answer #24*  **Image result for alh3 lewis** |  | *Answer #33*  **3CuBr2 + 2AlCl3 → 3CuCl2 +2AlBr3** |
| *Answer #25*  **2 mole TNT : 7 mole CO** |  |  |
| SET #5 **Questions #34 - 41** |  | SET #6 **Questions #42 - 49** |
| *Answer #34*  **0.103 moles** |  | *Answer #42*  **0.99g** |
| *Answer #35*  **1 mole Fe2(SO4)3 =  3 moles Na2SO4** |  | *Answer #43*  **8.2 mol** |
| *Answer #36*  **30 moles Na2SO4** |  | *Answer #44*  **1.23 x 1024 molecules** |
| *Answer #37*  **2H2 + O2 → 2H2O**  **0.94 moles water** |  | *Answer #45*  **5.44 x 10-5  mol B** |
| *Answer #38*   1. **Production of heat and light** 2. **Production of a gas** 3. **Formation of a precipitate** 4. **Change in color** |  | *Answer #46*  **Trigonal Planar** |
| *Answer #39*  **2Na + Cl2 → 2NaCl**  **1 mole Cl2** |  | *Answer #47*  **Valence  electrons** |
| *Answer #40*  **2NaCl+Ba→BaCl2+2Na**  **39.32 g Na** |  | *Answer #48*  **Gain 3 electrons** |
| *Answer #41*  **19.52L  F2** |  | *Answer #49*  **26 ve-** |
| SET #7 **Questions #50 - 58** |  | SET #8 **Questions #59 - 66** |
| *Answer #50*  **Ionic, covalent,  covalent, covalent,  ionic** |  | *Answer #59*  **1 lone pair** |
| *Answer #51*  **Mono, di, tri, tetra,  penta, hexa, hepta,  octa, nona, deca** |  | *Answer #60* |
| *Answer #52*  **Dicarbon  hexahhydride** |  | *Answer #61* |
| *Answer #53*  **Silver oxide** |  | *Answer #62*  **CHCl3 is more polar because greater electroneg. difference between atoms than in CHBr3** |
| *Answer #54*  **Copper (III) Nitrite** |  | *Answer #63*  **London forces** |
| *Answer #55*  **Sulfur hexoxide** |  | *Answer #64*  **Dipole-Dipole**  **(yes H is there, but no H-NOF:)** |
| *Answer #56* |  | *Answer #65*  **Network Covalent** |
| *Answer #57* |  | *Answer #66*  **4** |
| *Answer #58*  **0 lone pairs** |  |  |