|  |  |  |
| --- | --- | --- |
| SET #1**Questions #1 - 8** |  | SET #2**Questions #9 - 16** |
| *Answer #1***Strontium Oxide** |  | *Answer #9***Single replacement/single displacement** |
| *Answer #2***421.61 g/mol** |  | *Answer #10***CI2** |
| *Answer #3***Sea of electrons, delocalized electrons etc.** |  | *Answer #11* |
| *Answer #4***Covalent bond** |  | *Answer #12***tetrahedral** |
| *Answer #5***2 mol O2 : 1 mol CH4** |  | *Answer #13***68.15 g/mol** |
| *Answer #6***Hydrogen bonding****Remember H-NOF!** |  | *Answer #14***trigonal pyramidal** |
| *Answer #7***Metallic < Ionic Lattice < Network Covalent** |  | *Answer #15* |
| *Answer#8***Cu(SO4)2** |  | *Answer #16***Combustion!** |
| SET #3**Questions #17 - 25** |  | SET #4**Questions #26 - 33** |
| *Answer #17***AlPO4 + 3Rb(NO2) → Al(NO2)3 + Rb3(PO4)** |  | *Answer #26***Na2 CO3** |
| *Answer #18***Unequally!** |  | *Answer #27***106g/mol** |
| *Answer #19***10 mol ZnO!****Same molar ratio!** |  | *Answer #28***Fe2(SO4)3=400.1 g/mole** |
| *Answer #20***31.1g ZnO** |  | *Answer #29***0.111 moles Fe2(SO4)3** |
| *Answer #21***H2 F2 Cl2 Br2 I2 O2 N2** |  | *Answer #30***1087.7g KCl** |
| *Answer #22***77.98 g/mol** |  | *Answer #31***1592.68 gFe2(CO3)3** |
| *Answer #23***L of A→molA→mol B→L of B****76L H2O** |  | *Answer #32***Double displacement.** |
| *Answer #24***Image result for alh3 lewis** |  | *Answer #33***3CuBr2 + 2AlCl3 → 3CuCl2 +2AlBr3** |
| *Answer #25***2 mole TNT : 7 mole CO** |  |  |
| SET #5**Questions #34 - 41** |  | SET #6**Questions #42 - 49** |
| *Answer #34***0.103 moles** |  | *Answer #42***0.99g** |
| *Answer #35***1 mole Fe2(SO4)3 = 3 moles Na2SO4** |  | *Answer #43***8.2 mol** |
| *Answer #36***30 moles Na2SO4** |  | *Answer #44***1.23 x 1024 molecules** |
| *Answer #37***2H2 + O2 → 2H2O****0.94 moles water** |  | *Answer #45***5.44 x 10-5 mol B** |
| *Answer #38*1. **Production of heat and light**
2. **Production of a gas**
3. **Formation of a precipitate**
4. **Change in color**
 |  | *Answer #46***Trigonal Planar** |
| *Answer #39***2Na + Cl2 → 2NaCl****1 mole Cl2** |  | *Answer #47***Valence electrons** |
| *Answer #40***2NaCl+Ba→BaCl2+2Na****39.32 g Na** |  | *Answer #48***Gain 3 electrons** |
| *Answer #41***19.52L F2** |  | *Answer #49***26 ve-** |
| SET #7**Questions #50 - 58** |  | SET #8**Questions #59 - 66** |
| *Answer #50***Ionic, covalent, covalent, covalent, ionic** |  | *Answer #59***1 lone pair** |
| *Answer #51***Mono, di, tri, tetra, penta, hexa, hepta, octa, nona, deca** |  | *Answer #60* |
| *Answer #52***Dicarbon hexahhydride** |  | *Answer #61* |
| *Answer #53***Silver oxide** |  | *Answer #62***CHCl3 is more polar because greater electroneg. difference between atoms than in CHBr3** |
| *Answer #54***Copper (III) Nitrite** |  | *Answer #63***London forces** |
| *Answer #55***Sulfur hexoxide** |  | *Answer #64***Dipole-Dipole****(yes H is there, but no H-NOF:)** |
| *Answer #56* |  | *Answer #65***Network Covalent** |
| *Answer #57* |  | *Answer #66***4** |
| *Answer #58***0 lone pairs** |  |  |