**Name: Period: Seat#:**

**S-40**

You will use this worksheet to help record a summary of the major topics/ideas/facts for each chapter. Once you have spent some time in class remembering the main topics/ideas/facts for each chapter you will spend some time looking through your old worksheets to try and identify “representative questions” for each chapter – if your teacher could pick only a couple questions from each chapter to put on the final exam, which ones would your teacher pick? It is impossible to cover every single tiny bit of content on a final, so which questions from the year would assess if a student learned the most important aspects of the course?

* Participate in the class poster making activity.
* Use this worksheet during the “gallery walk” to capture the information on the posters.
	+ If you would like more space, you are welcome to use extra paper and staple it to this worksheet!
* Go through your old rainbow packets and identify questions on each worksheet that you think your teacher would consider putting on the final exam.
	+ Identify it by worksheet # and question #
* In class your teacher will show you the list they picked.
* Determine if you identified the same exact questions, similar questions, or if you were totally off the mark.
* You do not have to do these practice problems again unless you feel like it would be a good use of your self-study time.

|  |  |
| --- | --- |
| **Chapter 10****Thermochemistry** | **Chapter 9****Gas Laws** |
| **Chapter 12****Kinetics** | **Chapter 11****Solutions** |
| **Chapter 14****Acids and Bases** | **Chapter 13****Equilibrium** |
| **Poster 8****Things from 1st Semester that dont go away 2nd semester!**  | **Poster 7****Common mistakes, warnings, tips, etc** |

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| **Representative Questions Identification** |
| **Chapter #** | **WS #** | **My Choice** | **Teacher’s Choice of Representative Qs.** | **Q Matched/Similar** | **Off the Mark, Revisit WS** | **Comments** |
| **9** | 2 |  | 13 – 16, 18 |  |  | Any of the gas laws. |
| 3 |  | 5 -6, 8 |  |  | Ideal Gas Law, gases at STP, and molecular weight of gases |
| 5 |  | 8-9, 12 - 13 |  |  | Dalton’s Law, mole ratio of gases, collection of a gas over water |
| 8 |  | 1 - 3 |  |  | Gas stoichiometry  |

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| **Chapter #** | **WS #** | **My Choice** | **Teacher’s Choice of Representative Qs.** | **Q Matched/Similar** | **Off the Mark, Revisit WS** | **Comments** |
| **9***continued* | 9 |  | 2 - 3 |  |  | More gas stoichiometry – be able to combine with predicting products and balancing |
| 11 |  | 3 |  |  | Calculate molar mas of a gas when collected over water |
|  |
| **10** | 1 |  | 4, 9, 13 |  |  | Know vocab and examples |
| 2 |  | Any! |  |  | Use the q = mC∆T equation |
| 3 |  | 7-8 |  |  | Calorimetry practice |
| 6 |  | 4-13 |  |  | Know how to use heating and cooling curves to calculate energy gained/lost |
| 7 |  | 19, 21 |  |  | Molar heat calculations |
|  |
| **11** | 1 |  | 15 - 21 |  |  | Interpreting graphical data for solutions and vocabulary  |
| 2 |  | Any! |  |  | Calculate concentrations in M, ppm, g/L etc... |
| 3 |  | Any! |  |  | Calculate concentrations in M, ppm, g/L etc... |
| 4 |  | 1 - 21 |  |  | Calculate dilutions using M1V1=M2V2 |
| 5 |  | 1 - 13 |  |  | Calculate dilutions using M1V1=M2V2 |

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| --- | --- | --- | --- | --- | --- | --- |
| **Chapter #** | **WS #** | **My Choice** | **Teacher’s Choice of Representative Qs.** | **Q Matched/Similar** | **Off the Mark, Revisit WS** | **Comments** |
| **12** | 3 |  | 17, 20 |  |  | Calculating average rate from data, using rate expression |
| 4 |  | 6 - 10 |  |  | Calculating rate law and rate constant from data |
| 5 |  | 2, 7, 11 |  |  | Calculate instantaneous rate, rate law, and predict changes to rate based on changes in concentration of reactants |
|  |
| **13** | 2 |  | Any! |  |  | Predict changes to equilibrium using Le Chatlier’s Principle  |
| 5 |  | 11 - 18 |  |  | Write equilibrium expressions |
| 6 |  | 1 - 2 |  |  | Using ICE tables to calculate concentrations of reactants and products at equilibrium, predict shifts using Q compared to Keq |
| 7 |  | 11 - 14 |  |  | Using ICE tables to predict concentrations of reactants and products at equilibrium |
|  |
| **14** | 3 |  | 1 - 20 |  |  | Calculating pH, pOH, [H+], and/or [OH-] |
| 5 |  | 9 - 14 |  |  | Predicting if a solution is acidic, basic, or neutral based on pH or pOH calculations and using Kw  |
| 6 |  | Any! |  |  | Using ICE tables for calculating pH or pOH, conjugate pairs  |
| 8 |  | Any! |  |  | Predicting if a salt solution is acidic, basic, or neutral. |
| 9 |  | 1 -3  |  |  | Predicting if a salt solution is acidic, basic, or neutral. |
| 12 |  | 8 - 15 |  |  | Titration of acids/bases  |