**Name: Period: Seat#:**

**S-36**

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| 1. When calculating molarity, the volume needs to have what unit?
 | 1. The maximum amount of solute dissolved is called \_\_\_\_\_\_\_\_\_\_\_.
 | 1. Less than the maximum amount of solute dissolved is called \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
 |
| 1. More than the maximum amount of solute dissolved is called \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
 | 1. The solubility of solids goes \_\_\_\_\_\_\_\_\_\_ as the temperature is increased.
 | 1. The solubility of gases goes \_\_\_\_\_\_\_\_\_\_ as the temperature is increased.
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| 1. If you’re trying to make a diluted solution, you use the equation:
 | 1. When making a diluted solution the water added to the new solution is found by subtracting which two numbers?
 | 1. Factors that affect rate are:
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| 1. Factors that affect equilibrium position:
 | 1. Only \_\_\_\_\_\_\_\_\_ changes the equilibrium constant (keq)
 | 1. What is average rate?
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| 1. What is a rate expression? What is it used for?
 | 1. When you want the rate of one substance but you only have the rate for another substance, you can use the \_\_\_\_\_to solve for the missing rate. *Practice q: solve rate of h2 in terms of n2*
 | 1. The rate law only includes the concentrations of the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
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| 1. The equilibrium expression is \_\_\_\_\_\_\_\_\_\_\_ divided by \_\_\_\_\_\_\_\_\_\_\_
 | 1. The rate law exponents are called \_\_\_\_\_. Are they from the balanced equation coefficients or found experimentally?
 | 1. Are the exponents in an equilibrium expression from the balanced equation coefficients or found experimentally?
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| 1. Solids and liquids do or do not affect equilibrium?
 | 1. A large value for k indicates that the \_\_\_\_\_\_\_\_\_\_\_\_ side is favored and a small value for k indicates the \_\_\_\_\_\_\_\_\_\_\_ side is favored.
 | 1. $K'\_{eq}=$ ????
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| 1. If q is bigger than k, than the reaction will shift to the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
 | 1. If q is smaller than k, than the reaction will shift to the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
 | 1. I can use the 5% rule when:
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