**Name: Period: Seat#:**

**S-36**

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| 1. When calculating molarity, the volume needs to have what unit? | 1. The maximum amount of solute dissolved is called \_\_\_\_\_\_\_\_\_\_\_. | 1. Less than the maximum amount of solute dissolved is called \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_. |
| 1. More than the maximum amount of solute dissolved is called \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_. | 1. The solubility of solids goes \_\_\_\_\_\_\_\_\_\_ as the temperature is increased. | 1. The solubility of gases goes \_\_\_\_\_\_\_\_\_\_ as the temperature is increased. |
| 1. If you’re trying to make a diluted solution, you use the equation: | 1. When making a diluted solution the water added to the new solution is found by subtracting which two numbers? | 1. Factors that affect rate are: |
| 1. Factors that affect equilibrium position: | 1. Only \_\_\_\_\_\_\_\_\_ changes the equilibrium constant (keq) | 1. What is average rate? |
| 1. What is a rate expression?  What is it used for? | 1. When you want the rate of one substance but you only have the rate for another substance, you can use the \_\_\_\_\_to solve for the missing rate.  *Practice q: solve rate of h2 in terms of n2* | 1. The rate law only includes the concentrations of the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_. |
| 1. The equilibrium expression is \_\_\_\_\_\_\_\_\_\_\_ divided by \_\_\_\_\_\_\_\_\_\_\_ | 1. The rate law exponents are called \_\_\_\_\_. Are they from the balanced equation coefficients or found experimentally? | 1. Are the exponents in an equilibrium expression from the balanced equation coefficients or found experimentally? |
| 1. Solids and liquids do or do not affect equilibrium? | 1. A large value for k indicates that the \_\_\_\_\_\_\_\_\_\_\_\_ side is favored and a small value for k indicates the \_\_\_\_\_\_\_\_\_\_\_ side is favored. | 1. ???? |
| 1. If q is bigger than k, than the reaction will shift to the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_. | 1. If q is smaller than k, than the reaction will shift to the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_. | 1. I can use the 5% rule when: |
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