|  |  |
| --- | --- |
| **Dougherty Valley HS AP Chemistry** | **Name:** |
| **Equilibrium – Keq Determination** | **Date:** |
|  |  | **Period:** | **Seat #:** |

|  |
| --- |
| **Part I Data Table [Fill in title]:** |
| Temperature: \_\_\_\_\_\_\_\_\_ ºC |  |
| **Beaker** | **[FeSCN2+]** | **Absorbance** | **Work for [FeSCN2+] calculation: [to be included in the calculation section – new page in document]** |
| **1** |  |  |
| **2** |  |  |
| **3** |  |  |
| **4** |  |  |
| **Linear Regression equation**: |  |

|  |  |
| --- | --- |
| **Part II** |  |
| **Beaker** | **Absorbance** | **[FeSCN2+] at equilibrium** | **Work for [FeSCN2+] calculation: [to be included in the calculation section – new page in document]** |
| **A** |  |  |
| **B** |  |  |
| **C** |  |  |

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| Beaker A |   | **Fe3+** | **SCN–** | **FeSCN2+** |  **Part II Calculations [one calc per type for each beaker]** |
| Initial |   |   | 0.00 | **All calculations must be completed by hand**. Photo/scan your calculations and insert into another page in document below. Label it clearly |
| Change |   |   |   |
| Equilibrium |   |   |   |
|  |  |  |  |  |
| Beaker B |   | **Fe3+** | **SCN–** | **FeSCN2+** |
| Initial |  |  | 0.00 |
|  | Change |  |  |  |
|  | Equilibrium |  |  |  |
|  |  |  |  |  |
| Beaker C |   | **Fe3+** | **SCN–** | **FeSCN2+** |
| Initial |  |  | 0.00 |
|  | Change |   |   |   |
|  | Equilibrium |   |   |   |
|  |  |  |  |  |

1. Fill out the table below and then calculate the value of Keq for the reaction for each Beaker. Explain how you used the data to calculate Keq.

|  |  |  |  |
| --- | --- | --- | --- |
|   | **A** | **B** | **C** |
| [FeSCN2+] |  |  |  |
| [Fe3+] |  |  |  |
| [SCN–] |  |  |  |